

# **pH AND pOH**

Chapter 14 Section 14.2

The pH of a solution indicates how acidic or basic that solution is.

pH ranges of 0-7 = acidic

7 = neutral

7-14 = basic

Since  $\text{pH} + \text{pOH} = 14$  at 25°C (body temperature), then  $\text{pOH}$  can be calculated using either value.

$\text{pH} = 14 - \text{pOH}$

So if  $\text{pH} = 10.5$ ,  $\text{pOH} = 3.5$

$\text{pOH} = 14 - \text{pH}$

So if  $\text{pOH} = 10.5$ ,  $\text{pH} = 3.5$

Together  $\text{pH} + \text{pOH} = 14$

Complete the following chart.

	pH	pH	pOH	pOH	Acids or Bases
1.	normal	7	normal	7	Neutral
2.	$10^{-12}$	12	$10^{-2}$	2	acidic
3.	$10^{-10}$	10	normal	4	acidic
4.	normal	7	$10^{-7}$	7	neutral
5.	$10^{-7}$	7	$10^{-7}$	7	neutral
6.	$10^{-6}$	6	$10^{-8}$	8	basic
7.	$10^{-5}$	5	normal	9	basic
8.	normal	7	$10^{-7}$	7	neutral
9.	$10^{-3}$	3	$10^{-11}$	11	acidic
10.	$10^{-1}$	1	$10^{-13}$	13	acidic

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