

Electrochemistry Practice Test

Name:

Date:

Block:

I. Multiple Choice

1. When a metal undergoes corrosion, it:

- (a) loses electrons
(b) becomes reduced
(c) acts as an oxidizing agent
(d) decreases in oxidation number

2. At standard conditions, Fe^{+2} reacts spontaneously with

- (a) I_2
(b) Co
(c) Br^-
(d) Ag^+

3. Which of the following half-reactions is balanced?

- (a) $\text{SO}_4^{-2} + \text{H}_2\text{O} \rightleftharpoons \text{SO}_3^{-2} + 2\text{H}^+ + 2\text{e}^-$
(b) $\text{SO}_4^{-2} + \text{H}_2\text{O} + 2\text{e}^- \rightleftharpoons \text{SO}_3^{-2} + 2\text{H}^+$
(c) $\text{SO}_4^{-2} + 2\text{H}^+ + 2\text{e}^- \rightleftharpoons \text{SO}_3^{-2} + \text{H}_2\text{O}$
(d) $\text{SO}_4^{-2} + 2\text{H}^+ \rightleftharpoons \text{SO}_3^{-2} + \text{H}_2\text{O} + 2\text{e}^-$

4. During a redox reaction, the oxidizing agent:

- (a) reduces other species
(b) gains electrons
(c) increases in oxidation number
(d) becomes oxidized

5. For a given redox rxn, the oxidation # of tin changed from +2 to +4. As a result, tin:

- (a) lost 2 electrons and was reduced
(b) gained 2 electrons and was reduced
(c) lost 2 electrons and was oxidized
(d) gained 2 electrons and was oxidized

6. In which of the following compounds does carbon have an oxidation number of -2?

- (a) CO
(b) CO_2
(c) CH_2O
(d) CH_3OH

7. Which of the following equations represents a redox reaction?

- (a) $\text{ZnCl}_2(\text{aq}) \rightleftharpoons \text{Zn}^{+2}(\text{aq}) + 2\text{Cl}^-(\text{aq})$
(b) $\text{Zn}(\text{s}) + \text{Br}_2(\text{l}) \rightleftharpoons \text{ZnBr}_2(\text{s})$
(c) $\text{H}_2\text{CO}_3(\text{aq}) \rightleftharpoons \text{H}_2\text{O}(\text{l}) + \text{CO}_2(\text{g})$
(d) $2\text{NaI}(\text{aq}) + \text{Pb}(\text{NO}_3)_2(\text{aq}) \rightleftharpoons \text{PbI}_2(\text{s}) + 2\text{NaNO}_3(\text{aq})$

8. Consider the following reaction:

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Julia Schneider



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