

# CHEMISTRY 11 - FINAL EXAM STUDY GUIDE

## Unit 1: Safety in the Chemical Laboratory

- ensure that you understand all the basic safety rules to follow in the chemistry laboratory
- there will be a few safety questions on the final exam

## Unit 2: Introduction to Chemistry

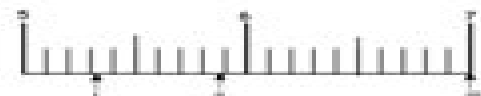
### UNIT CONVERSIONS

- you should know kilo, centi, milli and micro
- use conversion factors
- stop off at base unit (metre, litre, gram)
- always show your work
- always include the units
- (unknown) = (initial) X (conversion factor)
- DENSITY  $d = \frac{m}{V}$

### ACCURACY & PRECISION

- **significant figure** = a measured or meaningful digit; includes all certain digits and one uncertain
- all non-zero numbers significant
- leading zeros not significant
- trailing zeros significant if there is a decimal
- multiplication/division - round to lowest number of s.f.
- addition/subtraction - round to lowest number of decimal places
- hold all numbers in your calculator and round at the end
- **accurate** = close to correct or accepted value
- **precise** = more significant digits

Prefix	Abbreviation	Exponent
giga	G	$10^9$
mega	M	$10^6$
kilo	k	$10^3$
hecto	h	$10^2$
deca	da	$10^1$
deci	d	$10^{-1}$
centi	c	$10^{-2}$
milli	m	$10^{-3}$
micro	μ	$10^{-6}$
nano	n	$10^{-9}$
pico	p	$10^{-12}$



## Unit 3: The Physical Properties & Physical Changes of Substances

### DEFINITIONS

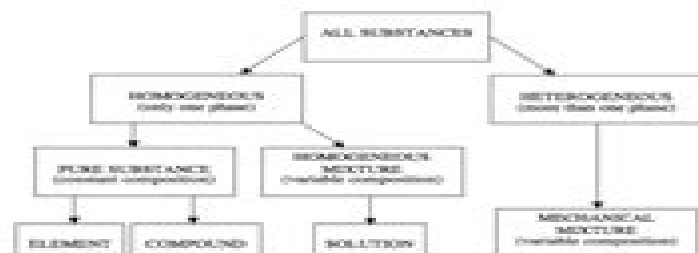
qualitative, quantitative, observation, interpretation, data, experiment, hypothesis, theory, law

### PROPERTIES OF MATTER

matter, physical property, chemical property, solid, liquid, gas, aqueous, malleability, ductility, lustre, viscosity.

### CLASSIFICATION OF MATTER

pure substance, mixture, homogeneous, heterogeneous, solution, solute, solvent, element, atom, molecule, ion, particle, compound



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**Diane Alford, Jennifer Hill**



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