

### Question 1

0 out of 1 points

A first order reaction  $A \rightarrow B + C$  has a rate constant of  $0.0125 \text{ s}^{-1}$ . If the initial concentration of  $A$  is  $0.500 \text{ M}$ , what is the concentration of  $A$  after 5 minutes? (Note: do not include units in your answer, do not use scientific notation, present your answer to 2 significant figures.)

Selected Answer: ☒ 25

Correct Answer: ☒ 0.0118

Answer range: 0.002 (0.0008 - 47- 0.0138)

### Question 2

0 out of 1 points

The reaction of nitrogen dioxide with fluorine generates nitryl fluoride:



To determine the rate law for this reaction, a chemist performed several initial-rate experiments using different initial concentration, with the details tabulated below.

Experiment (#)	$[\text{NO}_2] \text{ (M)}$	$[\text{F}_2] \text{ (M)}$	$[\text{NO}_2\text{F}] \text{ (M)}$	Initial Rate (M/s)
1	$1.0 \times 10^{-2}$	$5.0 \times 10^{-3}$	$0.10 \times 10^{-2}$	$2.0 \times 10^{-4}$
2	$2.0 \times 10^{-2}$	$5.0 \times 10^{-3}$	$0.10 \times 10^{-2}$	$4.0 \times 10^{-4}$
3	$2.0 \times 10^{-2}$	$5.0 \times 10^{-3}$	$1.00 \times 10^{-2}$	$4.0 \times 10^{-4}$
4	$2.0 \times 10^{-2}$	$7.5 \times 10^{-3}$	$0.10 \times 10^{-2}$	$6.0 \times 10^{-4}$

Determine the rate constant for this reaction. (Note: do not include units in your answer, do not use scientific notation, present your answer to 2 significant figures.)

Selected Answer: ☒ 120

Correct Answer: ☒ 40

Answer range: 2 (35.0 - 42.0) -47-

### Question 3

1 out of 1 points

Consider the following reaction with an equilibrium constant  $K_c = 20.06$ . Fill in the blanks with either less or greater.



From this information it can be conclude that the rate of the forward reaction is **[2] greater** than the rate of the reverse reaction and that the concentration of the reactants will be **[2] greater** than the concentration of the products.

Specified Answer: ☒ Great

# Chemistry 1020 Lab Manual Answers

**J. Philip Bromberg**



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