

Topic 4: Stoichiometry

- | | |
|---|---|
| 1) Chemical Formula
2) Chemical Equations
3) Relative Molecular/Formula mass (Mr) | 4) Moles (3 laws)
5) Empirical formula & Molecular Formula |
|---|---|

1) Chemical Formula

Valency or combining power:

- **The valency** of an element is the **number of electrons lost, gained or shared** to form compounds. The valency of an element depends on the number of **valence electrons**.
- In covalent molecules; the valency gives the number of covalent bonds which the atoms can form.
- In ionic compounds; the valency gives **the charge on the ions** of the element.
- The following table shows the valencies & charges of common elements used in chemistry:

Common metals	Element	Symbol	Valency	Charge	Common Non Metals	Element	Symbol	Valency	Charge
	Lithium	Li	1	+		Hydrogen	H	1	+1
	Sodium	Na	1	+		Chlorine	Cl	1	-1
	Potassium	K	1	+		Bromine	Br	1	-1
	☆ Silver	Ag	1	+		Iodine	I	1	-1
	Calcium	Ca	2	+2		Oxygen	O	2	-2
	Barium	Ba	2	+2		Sulfur	S	2	-2
	☆ Lead	Pb	2	+2		Nitrogen	N	3	-3
	Magnesium	Mg	2	+2		Phosphorus	P	3	-3
	Mercury	Hg	2	+2		Carbon	C	4	-4
	Zinc	Zn	2	+2		Silicon	Si	4	-4
	Aluminium	Al	3	+3	N.B.: • Element + Oxygen → Oxide • Element + Sulfur → Sulphide • Element + Nitrogen → Nitride • Element + Phosphorus → Phosphide • Element + Halogen → Halide • Element + Carbon → Carbide • Element + Hydrogen → Hydride				
	Iron II	Fe	2	+2					
	Iron III	Fe	3	+3					
	Copper I	Cu	1	+1					
	Copper II	Cu	2	+2					

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