13 Al -> "e + 26 Ma 48K -> = e + 40 G balance nuclear reaction equations aluminum-26 undergoes positron emission potassium-40 undergoes beta decay (electron) 238 () -> 4 He + 234 Th uranium-238 undergoes alpha decay electron configurations 152 25226 35236 451345 Write the electron configuration for C. Write the electron configuration for Cr. Write the shorthand notation for Bi. De ] 65244145 110603 NOMENCLATURE [naming compounds, writing formulas] Write the formula for each of the following compounds: a. Sulfur hexafluoride SFS b. Lithium nitride Link c. Chromium (III) carbonate Cr2 ((0))3 d. Tin (II) chloride Sn Ce2 e. Ammonium acetate NHUCZHOR (OF CHOCOONHY) f. Mercury (I) chloride Hazelz g. Potassium bromate KBrog h. Hydrosulfuric acid H2 Some i. Chloric acid Sulfurous acid 42300 Name each of the following compounds. a Cuso, copper (II) sulfate aluminum flooride b. AIF. c. HI d. NO narosen monovale c. H<sub>2</sub>Sc (40) hydroselen ic acid f. HNO, netric acid narous acrd g. HNO: NaHSO<sub>a</sub> Soulden

#### Half-life

Tritium (H-3) is a radioactive isotope of hydrogen with a half-life of 12.3 years. How long would it takes for a 40.0 g sample to decay down to 1.25 g? 40 -> 20 -> 10 -> 5 -> 2.5 -> 1.25 ->

Fe-61 has a half-life of 6.00 min. Of a 100.0 mg sample, how much will remain after 18.0 min?

3 half lives 100-750-25-3 [2.5mg]

#### Average atomic mass

The element copper has naturally occurring isotopes with mass numbers of 63 and 65,

The relative abundance and atomic masses are 69.2% for a mass of 63amu and 30.8% for a mass of 65amu. Calculate the average atomic mass of copper.

63 
$$\times .692 = 43.596 \rightarrow 49$$
  
65  $\times .309 = 20.02 \rightarrow 20.$   
 $\frac{+}{63.62} \rightarrow 64$ 

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**K Morrison** 

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#### **Table of Contents Chemistry Review Study Guide Answer Key**

- 1. Understanding the eBook Chemistry Review Study Guide Answer Key
  - The Rise of Digital Reading Chemistry Review Study Guide Answer Key
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Chemistry Review Study Guide Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemistry Review Study Guide Answer Key
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Chemistry Review Study Guide Answer Key
  - Personalized Recommendations
  - Chemistry Review Study Guide Answer Key User Reviews and Ratings
  - o Chemistry Review Study Guide Answer Key and Bestseller Lists
- 5. Accessing Chemistry Review Study Guide Answer Key Free and Paid eBooks

- Chemistry Review Study Guide Answer Key Public Domain eBooks
- Chemistry Review Study Guide Answer Key eBook Subscription Services
- Chemistry Review Study Guide Answer Key Budget-Friendly Options
- 6. Navigating Chemistry Review Study Guide Answer Key eBook Formats
  - o ePub, PDF, MOBI, and More
  - Chemistry Review Study Guide Answer Key Compatibility with Devices
  - Chemistry Review Study Guide Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemistry Review Study Guide Answer Key
  - Highlighting and Note-Taking Chemistry Review Study Guide Answer Key
  - Interactive Elements Chemistry Review Study Guide Answer Key
- 8. Staying Engaged with Chemistry Review Study Guide Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemistry Review Study Guide Answer Key
- 9. Balancing eBooks and Physical Books Chemistry Review Study Guide Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemistry Review Study Guide Answer Key
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Review Study Guide Answer Key
  - Setting Reading Goals Chemistry Review Study Guide Answer Key
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Review Study Guide Answer Key
  - Fact-Checking eBook Content of Chemistry Review Study Guide Answer Key
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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