

**Assessment #3 Review**

Name: \_\_\_\_\_

You must show all work for credit.

Due Date: \_\_\_\_\_ (No late packets will be accepted)

**I. Balance Equations***Balance the equations*

- $\text{Mg} + \underline{2} \text{H}_2\text{O} \rightarrow \text{Mg}(\text{OH})_2 + \text{H}_2$
- $\text{Mg} + \underline{2} \text{NaBr} \rightarrow \text{MgBr}_2 + \underline{2} \text{Na}$
- $\text{C}_3\text{H}_8 + \underline{5} \text{O}_2 \rightarrow \underline{3} \text{CO}_2 + \underline{4} \text{H}_2\text{O}$
- $\underline{4} \text{P} + \underline{5} \text{O}_2 \rightarrow \underline{2} \text{P}_2\text{O}_5$
- $\text{Fe}_2(\text{SO}_4)_3 + \underline{6} \text{KOH} \rightarrow \underline{3} \text{K}_2\text{SO}_4 + \underline{2} \text{Fe}(\text{OH})_3$
- $\text{Fe}_2\text{O}_3 + \underline{3} \text{H}_2 \rightarrow \text{Fe} + \underline{3} \text{H}_2\text{O}$

7. How many moles of hydrogen gas react with 1 mole of  $\text{Fe}_2\text{O}_3$  in reaction #6? 3

*Write the balanced equation for problems # 8 and 9:*

8. Hydrogen gas and nitrogen monoxide combine to form water and nitrogen gas



9. When sodium phosphate and calcium chloride are mixed they form calcium phosphate and sodium chloride



10. Using the following balanced equation, prove (or explain) the Law of Conservation of Mass.



2 N on left, 2 N on right, 6 H on left, 6 H on right, 2 on left, 2 on right.  
sub created, or lost

11. Using the balanced equation shown in problem #10, explain what the coefficients represent.

2 N, 6 H, 2 on left, 2 on right.  
sub created, or lost

12. Using the balanced equation shown in problem #10, list the products of the reaction.

2 NH<sub>3</sub>

13. Using the balanced equation shown in problem #10, list the reactants of the reaction.

N<sub>2</sub>, 3 H<sub>2</sub>

**II. Types of Reactions***Directions: Identify the type of reaction.*

S = synthesis

SD = single displacement

C = combustion

D = decomposition

DD = double displacement



# Chemistry Review Answers

**Alexander Stone Macnow**



## **Chemistry Review Answers:**

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## **Table of Contents Chemistry Review Answers**

1. Understanding the eBook Chemistry Review Answers
  - The Rise of Digital Reading Chemistry Review Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Review Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemistry Review Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Review Answers
  - Personalized Recommendations

- Chemistry Review Answers User Reviews and Ratings
- Chemistry Review Answers and Bestseller Lists
- 5. Accessing Chemistry Review Answers Free and Paid eBooks
  - Chemistry Review Answers Public Domain eBooks
  - Chemistry Review Answers eBook Subscription Services
  - Chemistry Review Answers Budget-Friendly Options
- 6. Navigating Chemistry Review Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Chemistry Review Answers Compatibility with Devices
  - Chemistry Review Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemistry Review Answers
  - Highlighting and Note-Taking Chemistry Review Answers
  - Interactive Elements Chemistry Review Answers
- 8. Staying Engaged with Chemistry Review Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemistry Review Answers
- 9. Balancing eBooks and Physical Books Chemistry Review Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemistry Review Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Review Answers
  - Setting Reading Goals Chemistry Review Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Review Answers
  - Fact-Checking eBook Content of Chemistry Review Answers

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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