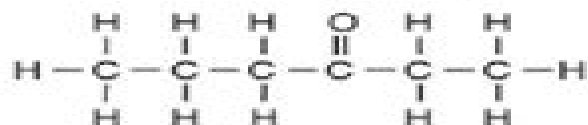


57. Given the formula representing a compound:



What is an IUPAC name for this compound?

- A. ethyl propanoate
- B. propyl ethanoate
- C. 3-hexanone
- D. 4-hexanone

58. As the temperature of a reaction increases, it is expected that the reacting particles collide

- A. more often and with greater force
- B. more often and with less force
- C. less often and with greater force
- D. less often and with less force

59. What is the total amount of heat required to completely melt 347 grams of ice at its melting point?

- A. 334 J
- B. 1450 J
- C. 116,000 J
- D. 784,000 J

60. Which two notations represent isotopes of the same element?

- A. ${}^9_7\text{N}$ and ${}^{10}_7\text{N}$
- B. ${}^{20}_7\text{N}$ and ${}^{20}_{10}\text{Ne}$
- C. ${}^9_7\text{N}$ and ${}^{10}_{10}\text{Ne}$
- D. ${}^{10}_7\text{N}$ and ${}^{10}_{10}\text{Ne}$

61. Which pair of atoms has the most polar bond?

- A. H-Br
- B. H-Cl
- C. I-Br
- D. I-Cl

62. What occurs when potassium reacts with chlorine to form potassium chloride?

- A. Electrons are shared and the bonding is ionic.
- B. Electrons are shared and the bonding is covalent.
- C. Electrons are transferred and the bonding is ionic.
- D. Electrons are transferred and the bonding is covalent.

63. Given the balanced equation representing a reaction:



What occurs as bonds are broken in one mole of H_2 molecules during this reaction?

- A. Energy is absorbed and one mole of unbonded hydrogen atoms is produced.
- B. Energy is absorbed and two moles of unbonded hydrogen atoms are produced.
- C. Energy is released and one mole of unbonded hydrogen atoms is produced.
- D. Energy is released and two moles of unbonded hydrogen atoms are produced.

64. Which equation shows conservation of charge?

- A. $\text{Cu} + \text{Ag}^+ \rightarrow \text{Cu}^{2+} + \text{Ag}$
- B. $\text{Mg} + \text{Zn}^{2+} \rightarrow 2\text{Mg}^{2+} + \text{Zn}$
- C. $2\text{F}_2 + \text{Br}^- \rightarrow 2\text{F}^- + \text{Br}_2$
- D. $2\text{I}^- + \text{Cl}_2 \rightarrow \text{I}_2 + 2\text{Cl}^-$

65. At STP, graphite and diamond are two solid forms of carbon. Which statement explains why these two forms of carbon differ in hardness?

- A. Graphite and diamond have different ionic radii.
- B. Graphite and diamond have different molecular structures.
- C. Graphite is a metal, but diamond is a nonmetal.
- D. Graphite is a good conductor of electricity, but diamond is a poor conductor of electricity.

66. Which electron configuration represents an excited state for an atom of calcium?

- A. 2-8-7-1
- B. 2-8-7-2
- C. 2-8-7-3
- D. 2-8-8-2

67. Which list of elements is arranged in order of increasing electronegativity?

- A. Be, Mg, Ca
- B. F, Cl, Br
- C. K, Ca, Sc
- D. Li, Na, K

68. Which radioisotope has the fastest rate of decay?

- A. ${}^{14}\text{C}$
- B. ${}^{40}\text{Ca}$
- C. ${}^{235}\text{Pu}$
- D. ${}^{40}\text{K}$

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