

"Section – B"**Marks: 32****Q-II** Write short answers of any eight (8) of the following parts. All parts carry equal marks.

1. Differentiate between compound and mixture.
2. What is electronic configuration?
3. How will you differentiate between representative and transition elements?
4. Why most atoms are chemically bonded to other atoms in nature?
5. Give an example of a non polar molecule with polar bonds. Give reason.
6. Why are the densities of gases lower than that of liquids?
7. Write factors affecting evaporation.
8. A bottle in a drug store contain a label "3 percent hydrogen peroxide". What does it mean?
9. Why NaOH is a strong but NH_4OH is weak electrolyte?
10. Write down some possible uses of an electrolytic cell.
11. Compare between metals and non metals.

Section "C"**Marks: 21****Note:** Answer any **THREE** questions. All questions carry equal marks.**Q.3** State the postulates of Neil Bohr's atomic theory.**Q.4** a. Distinguish between ionic and covalent bond.

b. Explain how ionization energies of elements vary across a period and a group in the periodic table.

Q.5 a. Define the following terms:

(i) Effusion (ii) Condensation (iii) Evaporation

b. Write a note on molarity.

Q.6 (a) Sketch the Daniel cell, labeling the cathode, anode and the direction of flow of the electrons.

(b) What happens during displacement reaction in halogens? Give at least one example.

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E Durkheim



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