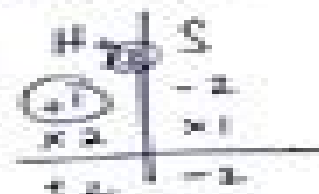
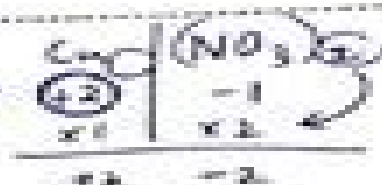


Worksheet #3: Naming Inorganic Compounds

1. Give the name for each of the following ionic compounds:

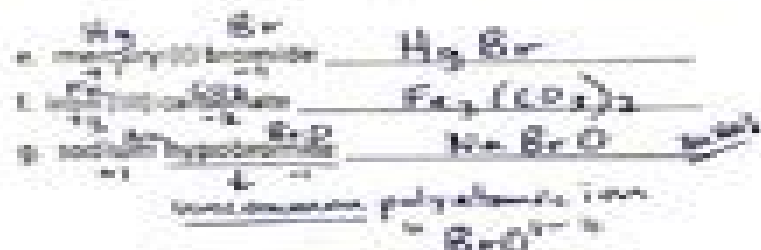
- AlF_3 Aluminum Fluoride
- Fe(OH)_2 Iron(II) Hydroxide
- $\text{Cu(NO}_3)_2$ Copper (II) Nitrate
- $\text{Ba(ClO}_3)_2$ Barium Perchlorate
- Li_3PO_4 Lithium Phosphate
- HgS Mercury (I) Sulfide
- $\text{Ca(CH}_3\text{COO)}_2$ Calcium Acetate
- $\text{Cr}_2\text{(CO}_3)_3$ Chromium (III) Carbonate
- $\text{K}_2\text{Cr}_2\text{O}_7$ Potassium Dichromate
- $\text{(NH}_4)_2\text{SO}_4$ Ammonium Sulfate



all of the underlined compounds are polyatomic ions → Memorize Them

2. Write the chemical formula for each of the following compounds:

- Copper(II) oxide Cu_2O
- potassium peroxide K_2O_2
- aluminum hydroxide Al(OH)_3
- Iron(III) bromide FeBr_3



3. Give the name or chemical formula, as appropriate, for each of the following acids:

- HNO_3 Nitric Acid
- HNO_2 Nitrous Acid
- H_3PO_4 Phosphoric Acid
phosphate PO_4^{3-} (the = oxy)

- hypochlorous acid HClO
 - iodic acid HIO_3
 - sulfurous acid H_2SO_3
- Band off of Iodate as IO_3^-

4. Give the name or chemical formula, as appropriate, for each of the following molecular substances:

- SF_6 Sulfur hexafluoride
- IF_5 Iodine pentafluoride
- XeO_3 Xenon Trioxide

- dinitrogen tetroxide N_2O_4
- hydrogen cyanide HCN
- tetraphosphorus hexasulfide P_4S_6

There are
all binary
ionic
compounds
(+/-)

3

Binary

3

8 - 14 e-
2 - 10 e-
10 - 18 e-
18 - 32 e-
32 - 50 e-
50 - 72 e-
72 - 86 e-
86 - 118 e-
118 - 132 e-
132 - 154 e-
154 - 176 e-
176 - 198 e-
198 - 210 e-
210 - 222 e-
222 - 234 e-
234 - 254 e-
254 - 274 e-
274 - 294 e-
294 - 314 e-
314 - 334 e-
334 - 354 e-
354 - 374 e-
374 - 394 e-
394 - 414 e-
414 - 434 e-
434 - 454 e-
454 - 474 e-
474 - 494 e-
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Chemistry Packet Answers

George Joseph Michell



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