

2

MATTER AND CHANGE

SECTION 2.1 PROPERTIES OF MATTER (pages 39–42)

This section helps you distinguish extensive from intensive properties and identify substances by their properties. It teaches you how to differentiate the three states of matter. It also defines a physical property and lists examples of physical properties and physical changes.

► Describing Matter (page 39)

1. The _____ mass _____ of an object is a measure of the amount of matter the object contains.

2. How does an extensive property differ from an intensive property?

An extensive property depends on the amount of matter, while an intensive property depends on the type of matter.

► Identifying Substances (page 40)

3. Matter that has a uniform and definite composition is called a _____ substance _____.

4. Is the following sentence true or false? All samples of a substance have different physical properties. _____ false _____

5. A physical property is a quality or condition of a substance that can be _____ observed _____ or _____ measured _____ without changing the substance's composition.

6. Circle the letter of the term that is NOT a physical property.

- a. hardness c. boiling point
b. color **d. melting**

7. Look at Table 2.1 on page 40. What is the melting point of bromine? _____ -7°C _____

8. Look at Table 2.1 on page 40. Circle the letter of the substance that is a yellow solid and melts at 115°C.

- a. sulfur**
b. chlorine
c. gold
d. copper

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Table of Contents Chemistry Matter And Change Outline

1. Understanding the eBook Chemistry Matter And Change Outline
 - The Rise of Digital Reading Chemistry Matter And Change Outline
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Matter And Change Outline
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Matter And Change Outline
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Matter And Change Outline

- Personalized Recommendations
- Chemistry Matter And Change Outline User Reviews and Ratings
- Chemistry Matter And Change Outline and Bestseller Lists
- 5. Accessing Chemistry Matter And Change Outline Free and Paid eBooks
 - Chemistry Matter And Change Outline Public Domain eBooks
 - Chemistry Matter And Change Outline eBook Subscription Services
 - Chemistry Matter And Change Outline Budget-Friendly Options
- 6. Navigating Chemistry Matter And Change Outline eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Matter And Change Outline Compatibility with Devices
 - Chemistry Matter And Change Outline Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Matter And Change Outline
 - Highlighting and Note-Taking Chemistry Matter And Change Outline
 - Interactive Elements Chemistry Matter And Change Outline
- 8. Staying Engaged with Chemistry Matter And Change Outline
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Matter And Change Outline
- 9. Balancing eBooks and Physical Books Chemistry Matter And Change Outline
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Matter And Change Outline
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Matter And Change Outline
 - Setting Reading Goals Chemistry Matter And Change Outline
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Matter And Change Outline

- Fact-Checking eBook Content of Chemistry Matter And Change Outline
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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