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## CHAPTER 2 THE COMPONENTS OF MATTER

### **FOLLOW-UP PROBLEMS**

- 2.1A **Plan:** An element has only one kind of atom; a compound is composed of at least two kinds of atoms. A mixture consists of two or more substances mixed together in the same container.  
**Solution:**  
(a) There is only one type of atom (blue) present, so this is an **element**.  
(b) Two different atoms (brown and green) appear in a fixed ratio of 1/1, so this is a **compound**.  
(c) These molecules consist of one type of atom (orange), so this is an **element**.
- 2.1B **Plan:** An element has only one kind of atom; a compound is composed of at least two kinds of atoms.  
**Solution:**  
The circle on the left contains molecules with either only orange atoms or only blue atoms. This is a **mixture of two different elements**. In the circle on the right, the molecules are composed of one orange atom and one blue atom so this is a **compound**.
- 2.2A **Plan:** Use the mass fraction of uranium in pitchblende (from Sample Problem 2.2) to find the mass of pitchblende that contains 2.3 t of uranium. Subtract the amount of uranium from that amount of pitchblende to obtain the mass of oxygen in that amount of pitchblende. Find the mass fraction of oxygen in pitchblende and multiply the amount of pitchblende by the mass fraction of oxygen to determine the mass of oxygen in the sample.  
**Solution:**  
Mass (t) of pitchblende = 2.3 t uranium  $\times \frac{84.2 \text{ t pitchblende}}{71.4 \text{ t uranium}} = 2.7123 \text{ t pitchblende}$   
Mass (t) of oxygen in 84.2 t of pitchblende = 84.2 t pitchblende – 71.4 t uranium = 12.8 t oxygen  
Mass (t) of oxygen = 2.7123 t pitchblende  $\times \frac{12.8 \text{ t oxygen}}{84.2 \text{ t pitchblende}} = 0.4123 \text{ t oxygen} = 0.41 \text{ t oxygen}$
- 2.2B **Plan:** Subtract the amount of silver from the amount of silver bromide to find the mass of bromine in 26.8 g of silver bromide. Use the mass fraction of silver in silver bromide to find the mass of silver in 3.57 g of silver bromide. Use the mass fraction of bromine in silver bromide to find the mass of bromine in 3.57 g of silver bromide.  
**Solution:**  
Mass (g) of bromine in 26.8 g silver bromide = 26.8 g silver bromide – 15.4 g silver = 11.4 g bromine  
Mass (g) of silver in 3.57 g silver bromide = 3.57 g silver bromide  $\times \frac{15.4 \text{ g silver}}{26.8 \text{ g silver bromide}} = 2.05 \text{ g silver}$

# Chemistry Matter And Change Assessment Solutions Manual

**Minjie Lin**



## **Chemistry Matter And Change Assessment Solutions Manual:**

Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment. New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections, assessment exercises at the end of each section and relevant chapter problems at the end of each chapter. A new Math Check allows quick access to the needed basic skill. The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem. Every concept in the text is clearly illustrated with one or more step-by-step examples. Making it Real essays have been updated to present timely and engaging real-world applications emphasizing the relevance of the material they are learning. This edition continues the end-of-chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter.

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