

1. Determine if a pure substance is made of atoms or molecules.

_____ C	_____ H ₂ O	_____ carbon dioxide	_____ sodium
_____ Ar	_____ Cu	_____ neon	_____ sulfur
_____ PBr ₃	_____ P ₄	_____ sulfur trioxide	_____ neon
_____ Ca	_____ N ₂	_____ silicon	_____ fluorine
_____ H ₂	_____ C ₆ H ₁₂ O ₆	_____ boron trioxide	_____ iron

2. Distinguish between pure substances and mixtures.

_____ NaCl & H ₂ O	_____ sand & sugar
_____ CO ₂ & O ₂	_____ gold
_____ CH ₄	_____ lead (II) nitrate dissolved in water
_____ 2Fe(C ₂ H ₃ O ₂) ₂ · 10H ₂ O	_____ nitrogen trihydride and oxygen
_____ Ca ²⁺ & SO ₄ ²⁻ & H ₂ O	_____ marbles and beads

3. Distinguish between elements and compounds.

_____ I ₂	_____ nitrogen monoxide
_____ Ag ₂ P	_____ nickel
_____ Sr ₃ N ₂	_____ chromium (III) sulfide
_____ CO	_____ aluminum bromide
_____ Cu	_____ fluorine

4. Distinguish between homogeneous mixtures (solutions) and heterogeneous mixtures.

- i. How can you tell if a mixture is homogeneous or heterogeneous?

5. Describe a chemical on the macroscopic, particle, and symbolic levels.

- i. Know how to represent a chemical with particle pictures.
ii. Know how to represent a chemical with symbols (formulas).

6. Know which elements are diatomic in nature and which are polyatomic in nature.

- i. Know the name and symbol of each diatomic (7) and polyatomic element (2) in nature.

7. Know the names and symbols of the first 54 elements on the periodic table.

- i. Be familiar with the first 54 elements so you can easily locate it on the periodic table.
ii. If you have them memorized, you will save time having to look it up on the periodic table.

8. Understand how to use the periodic table to identify groups/families of elements (and their chemical properties) known as alkali metals, alkaline earth metals, halogens, and noble gases.

- i. Which groups are alkali metals, alkaline earth metals, halogens, and noble gases found?

9. Distinguish between metals and nonmetals.

_____ H	_____ Cl	_____ lithium	_____ silver
_____ N	_____ Cr	_____ oxygen	_____ carbon
_____ Sr	_____ P	_____ arsenic	_____ sulfur

10. Understand the nature of the covalent bond and ionic bond.

- i. Covalent bonds are formed when two nonmetals fight over or share electrons.
ii. Ionic bonds are formed when a cation (being positively charged because it lost electrons) is attracted to an anion (being negatively charged because it gained electrons).
a. Ionic formulas are written in simplest whole number ratios.

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Brian Stewart



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Table of Contents Chemistry Ii Honors Study Guide

1. Understanding the eBook Chemistry Ii Honors Study Guide
 - The Rise of Digital Reading Chemistry Ii Honors Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Ii Honors Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Ii Honors Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Ii Honors Study Guide
 - Personalized Recommendations
 - Chemistry Ii Honors Study Guide User Reviews and Ratings
 - Chemistry Ii Honors Study Guide and Bestseller Lists

5. Accessing Chemistry II Honors Study Guide Free and Paid eBooks
 - Chemistry II Honors Study Guide Public Domain eBooks
 - Chemistry II Honors Study Guide eBook Subscription Services
 - Chemistry II Honors Study Guide Budget-Friendly Options
6. Navigating Chemistry II Honors Study Guide eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry II Honors Study Guide Compatibility with Devices
 - Chemistry II Honors Study Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry II Honors Study Guide
 - Highlighting and Note-Taking Chemistry II Honors Study Guide
 - Interactive Elements Chemistry II Honors Study Guide
8. Staying Engaged with Chemistry II Honors Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry II Honors Study Guide
9. Balancing eBooks and Physical Books Chemistry II Honors Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry II Honors Study Guide
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemistry II Honors Study Guide
 - Setting Reading Goals Chemistry II Honors Study Guide
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry II Honors Study Guide
 - Fact-Checking eBook Content of Chemistry II Honors Study Guide
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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