

# CHEM 121: Final Exam Study Guide

## Chapter 1

- Know the Scientific method
- Know the definitions for hypothesis, scientific law, and scientific theory.

## Chapter 2

- Length, mass, weight, volume
  - Know  $1\text{ cm}^3=1\text{ mL}$  and  $1\text{ dm}^3=1\text{ L}$
- Significant figures and Scientific notation
  - Rounding, in addition or subtraction, in multiplication or division
  - Keep as many sig figs until final answer
- Scientific notation
- Metric system
  - Know the metric prefixes: kilo (k), deci (d), centi (c), milli (m), and micro ( $\mu$ )
  - Be able to perform metric-metric conversions using these prefixes.
- Use the metric-English conversions provided ( $1\text{ in.}=2.54\text{ cm}$ ;  $1\text{ lb}=453.6\text{ g}$ ;  $1\text{ qt}=946\text{ mL}$ )
- Volume by calculation
  - $V_{\text{rectangular solid}} = \text{length} \times \text{width} \times \text{thickness}$
- Volume by displacement
- Density:  $d = \frac{m}{V}$ 
  - Calculate density, mass, or volume
  - Identify what items sink or float given densities of liquids and solids.
- Temperature
  - Know the formulas for converting "F-to-"C or "C-to-"F and K-to-"C or "C-to-K
- Percentage: ratio of parts per 100 parts
  - Given amount of part and whole, calculate %
  - Use a given % to solve for part or whole

## Chapter 3

- Know that matter is studied at the macroscopic, microscopic, particulate (molecular) levels
- Physical states of matter
  - Determine physical state of substances (solids, liquids, gases) given descriptions of volume, shape, particles moving, etc.
- Classify properties and changes as physical or chemical
  - Know terms for changes of state:
    - Melting, freezing, vaporizing, condensation, sublimation, deposition
- Classification of matter
  - Given examples, determine which are elements, compounds, or mixtures
  - Given molecular-level images, determine which are elements, compounds, or mixtures and solids, liquids, or gases
  - Distinguish between homogeneous and heterogeneous mixtures
- Chemical reaction:
  - reactants: starting materials
  - products: substances produced in reaction

# Chemistry Final Exam Study Guide

**Didier Musso**



## **Chemistry Final Exam Study Guide:**

**Hazmat Chemistry Study Guide (Second Edition)** Jill Meryl Levy, 2011      *AP Chemistry Exam Study Guide*

Accepted, 2019-09-19      *SAT Chemistry Test Prep Review--Exambusters Flash Cards* SAT II Exambusters, 2017-12-01 SAT CHEMISTRY Study Guide 700 questions and answers Essential definitions formulas concepts and sample problems Topics Introduction Matter Atoms Formulas Moles Reactions Elements Periodic Table Electrons Chemical Bonds Heat Gases Phase Changes Solutions Reaction Rates Equilibrium Acids and Bases Oxidation and Reduction Introduction to Organic Chemistry Radioactivity EXAMBUSTERS SAT II Prep Workbooks provide comprehensive SAT II review one fact at a time to prepare students to take practice SAT II tests Each SAT II study guide focuses on fundamental concepts and definitions a basic overview to begin studying for the SAT II exam Up to 600 questions and answers each volume in the SAT II series is a quick and easy focused read Reviewing SAT II flash cards is the first step toward more confident SAT II preparation and ultimately higher SAT II exam scores      Study Guide for Chemistry by Steven S. Zumdahl Martha B. Barrett, 1986      **Class 9**

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Chemistry for the IB Diploma Geoff Neuss, 2001 This concise guide provides the content needed for the Chemistry IB diploma at both Standard and Higher Level It follows the structure of the IB Programme exactly and includes all the options Each topic is presented on its own page for clarity Higher Level material is clearly indicated and there are plenty of practice questions The text is written with an awareness that English might not be the reader's first language

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