

**LOHS Honors Chemistry**  
**Balancing Chemical Equations**

**Part I: Balance these chemical equations. Record what type of reaction on the blank in the right margin.**

1.  $\text{BaCl}_2 + 2\text{KClO}_3 \rightarrow \text{Ba}(\text{ClO}_3)_2 + 2\text{KCl}$  DR
2.  $\text{Pb}(\text{OH})_2 \rightarrow \text{PbO} + \text{H}_2\text{O}$  D
3.  $2\text{Na} + 2\text{H}_2\text{O} \rightarrow 2\text{NaOH} + \text{H}_2$  SR
4.  $\text{Pb}(\text{NO}_3)_2 + \text{K}_2\text{CrO}_4 \rightarrow \text{PbCrO}_4 + 2\text{KNO}_3$  DR
5.  $\text{Al}_2(\text{SO}_4)_3 + 3\text{Ca}(\text{OH})_2 \rightarrow 2\text{Al}(\text{OH})_3 + 3\text{CaSO}_4$  DR
6.  $\text{Mg} + 2\text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2$  SR
7.  $\text{Sn} + 2\text{KOH} \rightarrow \text{K}_2\text{SnO}_2 + \text{H}_2$  SR
8.  $2\text{PaI}_5 \rightarrow 2\text{Pa} + 5\text{I}_2$  D
9.  $3\text{SiF}_4 + 3\text{H}_2\text{O} \rightarrow 2\text{H}_2\text{SiF}_6 + \text{H}_2\text{SiO}_3$  SR
10.  $\text{Cl}_2\text{O}_7 + \text{H}_2\text{O} \rightarrow 2\text{HClO}_4$  S
11.  $16\text{Na} + \text{S}_8 \rightarrow 8\text{Na}_2\text{S}$  S
12.  $\text{N}_2\text{O}_3 + \text{H}_2\text{O} \rightarrow 2\text{HNO}_2$  S

**Part II: Write balanced chemical equations for each of the reactions below. Assume that all ionic compounds are dissolved in water (aqueous). Also be sure include physical states where appropriate.**

1. Iron metal combines with oxygen gas to form iron (III) oxide solid



2. Chromium metal and sulfuric acid ( $\text{H}_2\text{SO}_4$ ) react to form chromium (III) sulfate solution and hydrogen gas.



3. Barium chromate solution and aqueous hydrochloric acid ( $\text{HCl}$ ) react to produce barium chloride solution and hydrogen chromate (solid).



# Chemistry Equations And Answers

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