

Chemical Reactions

Section 9.1 Reactions and Equations

pages 282–288

Practice Problems

pages 284–287

Write skeleton equations for the following word equations.

1. Hydrogen and bromine gases react to yield hydrogen bromide.

hydrogen(g) + bromine(g) →
hydrogen bromide(g)



2. When carbon monoxide and oxygen react, carbon dioxide forms.

carbon monoxide(g) + oxygen(g) →
carbon dioxide(g)



3. **Challenge** Write the word equation and the skeleton equation for the following reaction: when heated, solid potassium chlorate yields solid potassium chloride and oxygen gas.

potassium chlorate(s) →
potassium chloride(s) + oxygen(g)



Write chemical equations for each of the following reactions.

4. In water, iron(III) chloride reacts with sodium hydroxide, producing solid iron(III) hydroxide and sodium chloride.



5. Liquid carbon disulfide reacts with oxygen gas, producing carbon dioxide gas and sulfur dioxide gas.



6. **Challenge** A piece of zinc metal is added to a solution of hydrogen sulfate. This reaction produces a gas and a solution of zinc sulfate.



Section 9.1 Assessment

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7. **Explain** why it is important that a chemical equation be balanced.

Because mass is neither created nor destroyed in chemical reactions, the numbers of atoms of all elements must be equal on both sides of the reaction arrow.

8. **List** three types of physical evidence that indicate a chemical reaction has occurred.

Answers may include release or absorption of energy, change in color, change in odor, formation of a gas, or formation of a solid.

9. **Compare and contrast** a skeleton equation and a chemical equation.

The skeleton equation includes the formulas of reactants and products. The chemical equation gives the relative amounts of reactants and products.

10. **Explain** why it is important to reduce coefficients in a balanced equation to the lowest possible whole-number ratio.

Coefficients in the lowest ratio most clearly indicate the relative amounts of substances in a reaction.

11. **Analyze** When balancing a chemical equation, can you adjust the subscript in a formula? Explain.

No. Doing so changes the identity of the substance.

Chemistry Chemical Reactions Solutions Manual

**Julie Henderleiter, Charles Trapp, Peter
Atkins, Loretta Jones**



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Solutions Manual for Chemistry: Molecules Matter and Change, Fourth Edition Julie Henderleiter, Charles Trapp, Peter Atkins, Loretta Jones, 1999-08-13 This student companion is a supplement to Chemistry Molecules Matter and Change 4th edition with CD ROM It features guided reading strategies collaborative learning sheets and strategies for using CD ROM tools

Student Solutions Manual for Investigating Chemistry Jason Powell, 2008-12-09 Included here are step by step solutions with detailed explanations to the odd numbered questions and problems from the end of each chapter

The Practice of Chemistry Study Guide & Solutions Manual Pamela Mills, Amina El-Ashmawy, 2003-04-14 Designed to help students understand the material better and avoid common mistakes Also includes solutions and explanations to odd numbered exercises

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Physical Chemistry Student Solutions Manual Charles Trapp, Peter Atkins, Marshall Cady, 2006-08-11 Change 21

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Organic Chemistry Study Guide with Solutions Manual Neil E. Schore, 2007 The guide includes chapter introductions that highlight new material chapter outlines detailed comments for each chapter section a glossary and solutions to the end of chapter problems presented in a way that shows students how to reason their way to the answer

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Solutions Manual to Accompany Physical Chemistry for the Life Sciences C. A. Trapp, M. P. Cady, 2011 This solutions manual contains fully worked solutions to all end of chapter discussion questions and exercises featured in Physical Chemistry for the Life Sciences

Study Guide and Solutions Manual Neil E. Schore, 2002-08-02

Physical Chemistry, Solutions Manual Robert A. Alberty, Robert J. Silbey, 1996-08-20 This book provides thorough coverage of physical chemistry It demonstrates the power and limits of thermodynamics with a more systematic treatment of the second law and more focus on entropy It also covers current topics

in physical chemistry and shows how physical chemistry relates to daily life Includes many current applications such as lasers

Instructor's Solutions Manual to Accompany Atkins' Physical Chemistry, Ninth Edition C. A. Trapp, M. P. Cady, Carmen Giunta, 2010 The Instructor's solutions manual to accompany Atkins Physical Chemistry provides detailed solutions to the b exercises and the even numbered discussion questions and problems that feature in the ninth edition of Atkins Physical Chemistry The manual is intended for instructors and consists of material that is not available to undergraduates The manual is free to all adopters of the main text

Physical Chemistry for the Life Sciences Solutions Manual Maria Bohorquez, 2005-07-26 The Solutions Manual is a powerful study aid that contains the complete answers to all the exercises in the text These worked out solutions guide you through each step and help you refine your problem solving skills Used in conjunction with the text the Solutions Manual is one of the best ways to develop a fuller appreciation of chemical principles It can also be used to review material identify problem areas where more study is needed and test yourself before an exam Book jacket

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