

IQWST Chemical Bonding Study Guide 2021

-everything with chemical bonding revolves around valence electrons

-every atom (with the exception of hydrogen and helium) want to obtain 8 valence electrons in order to be stable → hydrogen and helium only need two valence electrons to fill their 1 electron shell.
Hydrogen will also tend to lose 1 valence electron so it will be stable.

-Lewis structure (electron dot diagrams) is used to visualize the valence electrons for atoms.

Ex.



- Ionic and metallic bonds involve a *transfer* of electrons
- Covalent bonds involve a *sharing* of electrons
 - o *Metallic bonds involve bonding between two metals*
 - o *Ionic bonding involves bonding between a metal and a non-metal*
 - o *Covalent bonding involves bonding between two non-metals*
- Lewis structure changes when you are showing a covalent bond between two atoms. Where the atoms are bonded, replace the electrons shared with dashes as shown below.



- Oxidation number is assigned to atoms that lose or gain electrons (ions). When ions are formed, an ionic bond or metallic bond will happen. Determining oxidation number helps to predict how atoms will interact and bond.

Ex. An Ionic bond between Sodium and Chlorine will happen. Sodium loses its 1 valence electron and Chlorine will take on that electron to fulfill its octet. When sodium loses its 1 valence electron it now has an octet. Oppositely charged ions attract and bond. Sodium has an oxidation number of 1⁺ and Chlorine has an oxidation number of 1⁻.



- Once atoms bond, you are left with a molecule or compound. It's important to determine the number of atoms in the formula.

Chemistry Chemical Bonding Study Guide Answers

Lei Shi



Chemistry Chemical Bonding Study Guide Answers:

Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment. New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections, assessment exercises at the end of each section and relevant chapter problems at the end of each chapter. A new Math Check allows quick access to the needed basic skill. The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem. Every concept in the text is clearly illustrated with one or more step-by-step examples. Making it Real essays have been updated to present timely and engaging real-world applications emphasizing the relevance of the material they are learning. This edition continues the end-of-chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter.

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