

# CHAPTER 15 STUDY GUIDE

## Energy and Chemical Change

### Section 15.1 Energy

*In your textbook, read about the nature of energy.*

In the space at the left, write *true* if the statement is true; if the statement is false, change the italicized word or phrase to make it true.

- \_\_\_\_\_ 1. *Energy* is the ability to do work or produce heat.
- \_\_\_\_\_ 2. The law of conservation of energy states that energy *can be* created and destroyed.
- \_\_\_\_\_ 3. Chemical potential energy is energy stored in a substance because of its *composition*.
- \_\_\_\_\_ 4. *Heat* is a form of energy that flows from a warmer object to a cooler object.
- \_\_\_\_\_ 5. A calorie is the amount of energy required to raise the temperature of *one gram* of pure water by one degree Celsius.
- \_\_\_\_\_ 6. A *calorie* is the SI unit of heat and energy.
- \_\_\_\_\_ 7. The *specific heat* of a substance is the amount of heat required to raise the temperature of one gram of that substance by one degree Celsius.
- \_\_\_\_\_ 8. *Kinetic energy* is energy of motion.
- \_\_\_\_\_ 9. Chemicals participating in a chemical reaction contain only *potential energy*.
- \_\_\_\_\_ 10. One nutritional Calorie is equal to *100 calories*.
- \_\_\_\_\_ 11. One calorie equals *4.184 joules*.
- \_\_\_\_\_ 12. When a fuel is burned, some of its *chemical potential energy* is lost as heat.
- \_\_\_\_\_ 13. To convert kilojoules to joules, *divide* the number of kilojoules by 1000 joules/1 kilojoule.

Answer the following question. Show all your work.

14. If the temperature of a 500.0-g sample of liquid water is raised 2.00°C, how much heat is absorbed by the water? The specific heat of liquid water is 4.184 J/(g·°C).

# Chemistry Chapter 15 Study Guide Answers

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**Chemistry, Student Study Guide** James E. Brady,Fred Senese,2008-01-28 The image on the front cover depicts a

carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane  $\text{CH}_4$  from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement

**Study Guide to Accompany Calculus for the Management, Life, and Social Sciences** Clyde Metz, 1984-01-01 Study Guide to Accompany Calculus for the Management Life and Social Sciences

General Chemistry Study Guide Sixth Edition Darrell D. Ebbing, 1999 Wastewater Operator Certification Study Guide John Giorgi, 2011-01-12

**CliffsNotes AP Chemistry** Bobrow Test Preparation Services, 2009-02-09 The book itself contains chapter length subject reviews on every subject tested on the AP Chemistry exam as well as both sample multiple choice and free response questions at each chapter s end Two full length practice tests with detailed answer explanations are included in the book

**Student Study Guide and Solutions Manual to accompany General Organic and Biological Chemistry, 1e** Kenneth W. Raymond, 2005-10-07

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**Hazmat Chemistry Study Guide (Second Edition)** Jill Meryl Levy, 2011

**A Study Guide for Audrey Niffenegger's "The Time Traveler's Wife"** Gale, Cengage Learning, 2015-03-13 A Study Guide for Audrey Niffenegger's *The Time Traveler's Wife* excerpted from Gale's acclaimed *Literary News For Students* This concise study guide includes plot summary character analysis author biography study questions historical context suggestions for further reading and much more For any literature project trust *Literary News For Students* for all of your research needs

**Chemistry** John Olmsted, Gregory M. Williams, 1997 Textbook outlining concepts of molecular science

*Study Guide and Solutions Book* Harold Hart, 1983

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### **Table of Contents Chemistry Chapter 15 Study Guide Answers**

1. Understanding the eBook Chemistry Chapter 15 Study Guide Answers
  - The Rise of Digital Reading Chemistry Chapter 15 Study Guide Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Chapter 15 Study Guide Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemistry Chapter 15 Study Guide Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Chapter 15 Study Guide Answers
  - Personalized Recommendations
  - Chemistry Chapter 15 Study Guide Answers User Reviews and Ratings



- Chemistry Chapter 15 Study Guide Answers and Bestseller Lists
- 5. Accessing Chemistry Chapter 15 Study Guide Answers Free and Paid eBooks
  - Chemistry Chapter 15 Study Guide Answers Public Domain eBooks
  - Chemistry Chapter 15 Study Guide Answers eBook Subscription Services
  - Chemistry Chapter 15 Study Guide Answers Budget-Friendly Options
- 6. Navigating Chemistry Chapter 15 Study Guide Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Chemistry Chapter 15 Study Guide Answers Compatibility with Devices
  - Chemistry Chapter 15 Study Guide Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemistry Chapter 15 Study Guide Answers
  - Highlighting and Note-Taking Chemistry Chapter 15 Study Guide Answers
  - Interactive Elements Chemistry Chapter 15 Study Guide Answers
- 8. Staying Engaged with Chemistry Chapter 15 Study Guide Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemistry Chapter 15 Study Guide Answers
- 9. Balancing eBooks and Physical Books Chemistry Chapter 15 Study Guide Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemistry Chapter 15 Study Guide Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Chapter 15 Study Guide Answers
  - Setting Reading Goals Chemistry Chapter 15 Study Guide Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Chapter 15 Study Guide Answers
  - Fact-Checking eBook Content of Chemistry Chapter 15 Study Guide Answers
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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