

## Chapter 1

A chemical change involves the change in composition while a physical change does not change in the composition of matter.

- 1.7 The reaction of calcium metal with water is a chemical change resulting in the formation of new compounds, hydrogen gas and calcium hydroxide. It is not stated in the problem, but the water also increases in temperature, which is a physical change.
- 1.8 These are physical changes.
- 1.9
- (a) An element is a pure substance that cannot be decomposed into something simpler.
  - (b) A compound is a pure substance that is composed of two or more elements in some fixed or characteristic proportion.
  - (c) Mixtures result from combinations of pure substances in varying proportions.
  - (d) A homogeneous mixture has one phase. It has the same properties throughout the sample.
  - (e) A heterogeneous mixture has more than one phase. The different phases have different properties.
  - (f) A phase is a region of a mixture that has properties that are different from other regions of the mixture.
  - (g) A solution is a homogeneous mixture.
- 1.10 Changing a compound into its element is a chemical change.
- 1.11
- |        |        |        |        |
|--------|--------|--------|--------|
| (a) Cl | (b) S  | (c) Fe | (d) Ag |
| (e) Na | (f) P  | (g) I  | (h) Cu |
| (i) Hg | (j) Ca |        |        |
- 1.12
- |               |               |             |              |
|---------------|---------------|-------------|--------------|
| (a) potassium | (b) zinc      | (c) silicon | (d) tin      |
| (e) manganese | (f) magnesium | (g) nickel  | (h) aluminum |
| (i) carbon    | (j) nitrogen  |             |              |
- 1.13
- (a) This is a heterogeneous mixture.
  - (b) This is a pure substance and is an element, such as  $H_2$ ,  $O_2$ ,  $N_2$  or a halogen.
  - (c) This is a homogeneous mixture.
  - (d) This is a pure substance and is a molecule such as  $H_2O$ .
- 1.14
- (a) Diagrams (a) and (d) contain pure elements.
  - (b) Diagram (c) contains a compound.
  - (c) Diagram (a) and (b) contain diatomic molecules.
- 1.15 The first law of chemical combination is the law of conservation of mass: no detectable gain or loss of mass occurs in chemical reactions. The other law is the law of definite proportions: in a given chemical compound, the elements are always combined in the same proportions by mass.
- 1.16 Conservation of mass derives from the postulate that atoms are not destroyed in normal chemical reactions. The Law of Definite Proportions derives from the notion that compound substances are always composed of the same types and numbers of atoms of the various elements in the compound.
- 1.17 This is the Law of Definite Proportions, which guarantees that a single pure substance is always composed of the same ratio of masses of the elements that compose it.
- 1.18 The law of multiple proportions states that when elements combine to make a molecule the ratio of atoms is always a small, whole number. When elements combine to form more than one compound the ratio of one element in the compound compared to a fixed amount of a second element is always a small, whole number.

The first compound in the diagram has two blue and three red atoms. The second compound has two blue and one red atom. Since each compound has two blue atoms we can compare the ratio of red atoms between the

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**Lei Shi**



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