

# ORGO ACIDS & BASES

## CHEAT SHEET

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Arrhenius

Acid = Litmus Red  
 $H^+$  in  $H_2O$

Base Litmus Blue  
 $OH^-$  in  $H_2O$

Bronsted-Lowry

$H^+$  donor  
 ex.  $HCl$   
 ex.  $NH_4Cl$

Accepts  $H^+$   
 ex.  $NaOH$   
 ex.  $NH_3$

Lewis (electrons)

Accepts  $e^-$  pair  
 ex.  $AlCl_3$

$e^-$  pair donor  
 ex.  $Cl^-$

Acid - Base Reaction



Equations to recognize

$$K_a = \frac{[H^+][A^-]}{[HA]} \quad pK_a = -\log(K_a) \rightarrow K_a \propto [H^+] \propto \frac{1}{pK_a}$$

Strong Acid

$\uparrow [H^+]$

$\uparrow K_a$

$\downarrow pK_a$

Weak Acid

$\downarrow [H^+]$

$\downarrow K_a$

$\uparrow pK_a$

$$pH = -\log [H^+] \\ pOH = -\log [OH^-]$$

$$pH + pOH = 14 \\ pK_a + pK_b = pK_w = 14$$

pH scale  
 $pH \rightarrow$   
 $\leftarrow pOH$

$pH = 1$   
 Acidic  
 $pOH = 14$

$pH = 7$   
 Neutral  
 $pOH = 7$

$pH = 14$   
 Basic  
 $pOH = 1$

STRONG ACIDS FORM STABLE CONJUGATE BASES

CARIO

C = charge of acid or conj. base

A = Atom holding charge

R = Resonance

I = Inductive effect

O = orbital / hybridization

See videos for full explanation + examples [LEAH4SCI.COM/acidbase](http://LEAH4SCI.COM/acidbase)

Charge:  $+ / \delta^+$  more acidic } when compare species  
 $- / \delta^-$  more basic }

Atom: In same period  $\uparrow$  eneg  $\uparrow$  acidity  
 In same group  $\uparrow$  size  $\uparrow$  acidity



Resonance:  $\uparrow$  Res  $\uparrow$  charge distribution  
 $\uparrow$  acidity (Aromaticity =  $\uparrow$  stable)

Inductive Effect:  $\uparrow$  e-neg nearby atom  
 $\uparrow$  acidity

Orbital / hybridization  $\uparrow$  % s  $\uparrow$  acidity  
 $sp = 50\% > sp^2 = 33\% > sp^3 = 25\%$

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# Chemistry Acids And Bases Study Guide

**Norman E. Griswold, Henry Fuller  
Holtzclaw**



## Chemistry Acids And Bases Study Guide:

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The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment. New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections, assessment exercises at the end of each section, and relevant chapter problems at the end of each chapter. A new Math Check allows quick access to the needed basic skill. The first chapter now includes brief introductions to several fundamental chemical concepts, and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem. Every concept in the text is clearly illustrated with one or more step-by-step examples. Making it Real essays have been updated to present timely and engaging real-world applications, emphasizing the relevance of the material they are learning. This edition continues the end-of-chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter. Chemistry, Student Study Guide James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst. Carbon nanotubes are a recently discovered allotrope of carbon. Three other allotropes of carbon, buckyballs, graphite, and diamond, are illustrated at the left, as is the molecule methane,  $\text{CH}_4$ , from which nanotubes and buckyballs can be made. The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet. The study of chemistry also follows from the simple to the more complex, and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement. The Practice of Chemistry Study Guide & Solutions Manual Pamela Mills, Amina El-Ashmawy, 2003-04-14 Designed to help students understand the material better and avoid common mistakes. Also includes solutions and explanations to odd-numbered exercises. **CHEMISTRY HANDBOOK & STUDY GUIDE Gr11-12 NE** Kevin Smith, 2024-02-01 A comprehensive summary of Grade 11/12 Physics. Simple logical summaries with example exam questions and work-through solutions. The book covers the fundamentals of Grade 11/12 Physics and complements the material in any class text. *Organic Chemistry Study Guide with Solutions Manual* Neil E. Schore, 2007 The guide includes chapter introductions that highlight new material, chapter outlines, detailed comments for each chapter section, a glossary, and solutions to the end-of-chapter problems presented in a way that shows students how to reason their way to the answer. **Organic Chemistry, 12e Binder Ready Version Study Guide & Student Solutions Manual** T. W. Graham Solomons, Craig B. Fryhle, Scott A. Snyder, 2016-04-11 This is the Student Study Guide Solutions Manual to accompany Organic Chemistry 12th Edition. The 12th edition of Organic Chemistry continues Solomons, Fryhle, Snyder's tradition of excellence in teaching and preparing students for success in the

organic classroom and beyond A central theme of the authors approach to organic chemistry is to emphasize the relationship between structure and reactivity To accomplish this the content is organized in a way that combines the most useful features of a functional group approach with one largely based on reaction mechanisms The authors philosophy is to emphasize mechanisms and their common aspects as often as possible and at the same time use the unifying features of functional groups as the basis for most chapters The structural aspects of the authors approach show students what organic chemistry is Mechanistic aspects of their approach show students how it works And wherever an opportunity arises the authors show students what it does in living systems and the physical world around us

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