

Academic Chemistry 2nd Semester Final 2022

43. A solution in which the $[\text{OH}^-] = 1.0 \times 10^{-2} \text{ M}$ is:
a. acidic b. basic c. neutral d. none of the above
44. If the pH of a solution is 3.0, what is the $[\text{OH}^-]$?
a. 3 M b. 11 M c. $1 \times 10^{-3} \text{ M}$ d. $1 \times 10^{-11} \text{ M}$
45. If a solution has a pOH of 10.35 it is a(n):
a. acid b. base c. neutral solution
46. What is the pOH of a 0.55 M solution of HF?
a. 14 b. 3.70 c. 1.70 d. 3.41
47. What is the pH of a 0.19 M solution of HNO_3 ?
a. 0.72 b. 2.14 c. 4.88 d. 1.90

Chemistry 2nd Semester Finals

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