

## Chemistry Semester 1 Final Exam Review

1. What is a definition of Chemistry?  
**Study of matter**
  2. Matter includes all of the following EXCEPT:  
a. air      **b. light**      c. smoke      d. water vapor
  3. What is matter?  
**Something that has mass and takes up space (volume)**
  4. A physical change occurs when a  
a. peach spoils      c. copper bowl tarnishes  
b. bracelet turns your wrist green      **d. glue gun melts a glue stick**
  5. The particles in a solid are  
**a. packed closely together**      c. constantly in motion  
b. very far apart      d. able to slide past one another
- Of the following groups, 1, 2, 14, 15, 17, 18:
6. Group(s) **1, 2** contains only metals
  7. Group (s) **18** is called a Noble gases.
  8. Group(s) **17** has 7 electrons on the valence shell
  9. Group(s) **15** needs to gain 3 electrons to fill a stable octet
  10. Group(s) **2** has a 2+ oxidation state
  11. Group(s) **2** has an electron dot notation of  $\bullet X \bullet$
  12. Using the periodic table, an element similar to carbon would be **14**
  13. Based on their location on the periodic table, you could infer that **18/noble gases** are very unreactive
  14. The horizontal row on the periodic table is called a(n) **period**
  15. The vertical column of blocks on the periodic table is called a(n) **group/family**
  16. The elements on the zigzag line in the periodic table are **metalloids**
  17. Which is NOT a property of metal?  
a. malleability      c. **unreactivity**  
b. ability to lose electrons      d. ability to conduct heat
  18. A volume of 1 cc ( $\text{cm}^3$ ) is the same as **1** mL.
  19. To determine density, the quantities that must be measured are **volume** and **mass**
  20. The relationship between the mass  $m$  of a material, its volume  $V$ , and its density  $D$  is  
a.  $V = mD$       b.  $V/m = D$       c.  **$DV = m$**       d.  $D + V = m$
  28. The density of aluminum is  $2.70 \text{ g/cm}^3$ . The volume of a solid piece of Aluminum is  $1.50 \text{ cm}^3$ . Find the mass **4.05 grams**
  21. The density of pure diamond is  $3.5 \text{ g/cm}^3$ . The mass of a diamond is 0.25 g. Find the volume= **0.71  $\text{cm}^3$**
  22. What statement about density is true?  
a. two samples of a pure substance may have different densities if they are different sizes  
b. the density of a sample depends on its location on Earth  
c. a cylinder is always used to measure the volume  
**d. density is a physical property that remains constant for a pure substance**
  23. The number of grams equal to 0.5 kg is **500 g**
  24. In a graph the **independent** variable is placed on the x axis and the **dependant** variable is placed on the y axis.
  25. The variable that you change in the lab is the **independent** variable. The variable that responds to the change is called the **dependant** variable. All other things do not change and are called **constants**
  26. At first Dalton thought the atom was solid, spherical, **indivisible**, then Rutherford discovered the **nucleus** of the atom, then the subatomic particles **protons**, **neutrons**, and **electrons** were discovered
  27. The positively charged particle is the **proton** found in the **nucleus** of the atom
  28. The subatomic particle that has about the same mass as the proton, but with no electrical charge, is called a **neutron** and is found in the **nucleus** of the atom
  29. The **electron** is negatively charged and is found in the **electron cloud/outer shell** of the atom
  30. **Bohr** is the person credited with placing electrons on levels and **Schrodinger** established the electron cloud theory.
  31. The forces that hold the particles together in the nucleus are **columbic /electrostatic** forces
  32. Isotopes are atoms of the same element that **have different number of neutrons**
  33. All isotopes of hydrogen contain **1 proton & 1 electron**
  34. What is the atomic number for Nitrogen? **7** To what subatomic particle does this correspond? **# of protons**
  35. Zinc -66 has **30** protons and **36** neutrons
  36. What is the electron configuration for sodium?  **$[\text{Ne}]3s^1$**  Chlorine?  **$[\text{Ne}]3s^23p^5$**  Neon?  **$[\text{He}]2s^22p^6$**
  37. What is the dot notation for magnesium? Bromine? Argon?

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**Jenny Harrison**



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