

ORGO ACIDS & BASES CHEAT SHEET

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Arrhenius	Acid - Litmus Red H^+ in H_2O ex. HCl	Base Litmus Blue OH^- in H_2O ex. $NaOH$
Bronsted-Lowry	H^+ donor ex. NH_4Cl	Accepts H^+ ex. NH_3
Lewis (reactions)	Accepts e^- pair ex. $AlCl_3$	e^- pair donor ex. Cl^-

Acid - Base Reaction



Equations to recognize

$$K_a = \frac{[H^+][A^-]}{[HA]} \quad pK_a = -\log(K_a) \rightarrow K_a \propto [H^+] \propto \frac{1}{pK_a}$$

Strong Acid	$\uparrow [H^+]$	$\uparrow K_a$	$\downarrow pK_a$
Weak Acid	$\downarrow [H^+]$	$\downarrow K_a$	$\uparrow pK_a$

$$pH = -\log \frac{[H^+]}{[OH^-]} \quad pOH = -\log [OH^-]$$

$$pH + pOH = 14 \quad pK_a + pK_b = pK_w = 14$$

pH scale
 \leftarrow pH
pOH \rightarrow

pH = 1
Acidic
pOH = 14

pH = 7
Neutral
pOH = 7

Basic
pH = 14
pOH = 1

STRONG ACIDS FORM STABLE CONJUGATE BASES

CARIO

- C = charge of acid or conj. base
- A = Atom holding charge
- R = Resonance
- I = Inductive effect
- O = orbital / hybridization

See videos for full explanation + examples LEAH4SCI.COM/acidbase

Charge: $+ / \delta^+$ more acidic } when compare species
 $- / \delta^-$ more basic }

Atom: In same period \uparrow eneg \uparrow acidity
In same group \uparrow size \uparrow acidity



Resonance: \uparrow Res \uparrow charge distribution
 \uparrow acidity (Aromaticity = \uparrow stable)

Inductive Effect: \uparrow e-neg nearby atom
 \uparrow acidity

Orbital / hybridization \uparrow % s \uparrow acidity
 $sp = 50\% > sp^2 = 33\% > sp^3 = 25\%$

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Chemistry 01 Study Guide

Xiaolong Qi



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Table of Contents Chemistry 01 Study Guide

1. Understanding the eBook Chemistry 01 Study Guide
 - The Rise of Digital Reading Chemistry 01 Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry 01 Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry 01 Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry 01 Study Guide
 - Personalized Recommendations
 - Chemistry 01 Study Guide User Reviews and Ratings
 - Chemistry 01 Study Guide and Bestseller Lists
5. Accessing Chemistry 01 Study Guide Free and Paid eBooks
 - Chemistry 01 Study Guide Public Domain eBooks
 - Chemistry 01 Study Guide eBook Subscription Services
 - Chemistry 01 Study Guide Budget-Friendly Options

6. Navigating Chemistry 01 Study Guide eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry 01 Study Guide Compatibility with Devices
 - Chemistry 01 Study Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry 01 Study Guide
 - Highlighting and Note-Taking Chemistry 01 Study Guide
 - Interactive Elements Chemistry 01 Study Guide
8. Staying Engaged with Chemistry 01 Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry 01 Study Guide
9. Balancing eBooks and Physical Books Chemistry 01 Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry 01 Study Guide
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemistry 01 Study Guide
 - Setting Reading Goals Chemistry 01 Study Guide
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry 01 Study Guide
 - Fact-Checking eBook Content of Chemistry 01 Study Guide
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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