

## CHEMICAL QUANTITIES STUDY GUIDE

### Atomic Mass

- The mass of a single atom of an element (amu)
- Mass of one mole of the atom = atomic mass x amu
- $1 \text{ amu} = 1.66 \times 10^{-24} \text{ g}$

### Avogadro's Number

- $6.022 \times 10^{23}$
- Six hundred 2 sextillion two hundred quintillion
- Represents the number of atoms that are in 1 mole of a substance

### Molar Mass

- The sum of the total mass in grams of all the atoms that make up a mole of a particular molecule
- Molar mass of an element = number underneath name on periodic table
  - For a compound, calculate molar mass of each element and add

### Mole Conversions

- Moles = mass/molar mass

$$\begin{array}{ccccc} \text{MASS (g)} & \begin{array}{c} \times \\ \hline \text{MOLAR MASS} \\ \hline + \end{array} & \boxed{\text{MOLE}} & \begin{array}{c} \times \\ \hline 6.022 \times 10^{23} \\ \hline + \end{array} & \text{FORMULA} \\ & & \begin{array}{c} \times \quad | \quad 22.4 \quad | \quad + \\ \hline \text{VOLUME} \end{array} & & \text{UNITS} \end{array}$$

# Chemical Quantities Study Guide

**Daniela Niemeyer**



## Chemical Quantities Study Guide:

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