

Consider liquid water at 25°C and 1 atm. How many times does the volume increase if the water is vaporized and vapor at 100°C and 1 atm?

From the steam tables at 25°C and 1 atm,  $v_1 = 1.000 \text{ cm}^3/\text{gm}$ . At 100°C and 1 atm,  $v_2 = 1673.0 \text{ cm}^3/\text{gm}$ . Thus, the volume expansion ratio between the liquid and the vapor is 1673 times. Liquid water exposed to hot oil will flash explosively, leading to considerable damage due to the expansion of the liquid.

(a) 
$$Q_{ave} = \dot{V} = \frac{\pi r^2}{4} \cdot \frac{(\Delta p)}{4} = \frac{\pi (0.4)^2}{4} \cdot \frac{(30)}{4} = 706.5 \text{ ft}^3$$
$$= (706.5)(1.44) = 1.017 \times 10^5 \text{ in}^3$$
$$P = F/A = \frac{200 \text{ lb}}{1.017 \times 10^5 \text{ in}^2} = 1.97 \times 10^{-3} \text{ psia}$$

Since  $1 \text{ atm} = 14.7 \text{ psia} = 33.9 \text{ ft of H}_2\text{O}$   
 $= 40.9 \text{ in of water}$

Therefore

$$P = (1.97 \times 10^{-2} \text{ Pa}) \left( \frac{406.9 \text{ J}}{14.7 \text{ Pa} \cdot \text{m}^3} \right)$$

$P = 0.034$  in. of water (40+ inch)

# Chemical Process Safety 3rd Edition Solution Manual

**Clifford Lane**



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