

15. The term **polar covalent** refers to bonds where the electrons are not equally shared—there is a preference for one nucleus over the other. Discuss with your group what region of the Bond-Type Triangle in Model 4 would contain substances with polar covalent bonding.

Polar covalent bond happens mainly within ionic and covalent areas on the triangle.

16. The term **semimetal** refers to metallic bonds where the electrons are more tightly held. These substances will exhibit fewer metallic properties and more covalent properties. Discuss with your group what region of the Bond-Type Triangle in Model 4 would contain semimetallic substances.

Semimetals are within metallic and covalent areas on the triangle.

17. Use the Bond-Type Triangle in Model 4 to determine the type of bonding of these substances. Use ionic, covalent, metallic, polar covalent and semimetallic as your answers.

Substance	Electronegativity (EN)		AvgEN	Δ EN	Type of Bonding
	First atom	Second atom			
ICl (s)	2.5	3.0	2.75	0.5	polar cov.
MgO (s)	1.2	3.5	2.35	2.3	ion.
GaAs (s)	1.6	2.0	1.8	0.4	semim.
Si (s)	1.8	1.8	1.8	0	semim.
Cd ₃ Mg (s)	1.7	1.2	1.45	0.5	metallic
P ₄ O ₇ (s)	2.1	3.5	2.8	1.4	polar cov.
NCl ₃ (l)	3.0	3.0	3.0	0	covalent

Chemical Bonding Pogil Activity 5

Margaret Irene Lindsay



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