

CHEMFILE MINI-GUIDE TO PROBLEM SOLVING

Gas Laws

Chemists found that there were relationships among temperature, volume, pressure, and quantity of a gas that could be described mathematically. This chapter deals with Boyle's law, Charles's law, Gay-Lussac's law, the combined gas law, and Dalton's law of partial pressures. These laws have one condition in common. They all assume that the molar amount of gas does not change. In other words, these laws work correctly only when no additional gas enters a system and when no gas leaks out of it. Remember also that a law describes a fact of nature. Gases do not "obey" laws. The law does not dictate the behavior of the gas. Rather, each gas law describes a certain behavior of gas that occurs if conditions are right.

BOYLE'S LAW

Robert Boyle, a British chemist who lived from 1627 to 1691 formulated the first gas law, now known as Boyle's law. This law describes the relationship between the pressure and volume of a sample of gas confined in a container. Boyle found that gases compress, much like a spring, when the pressure on the gas is increased. He also found that they "spring back" when the pressure is lowered. By "springing back" he meant that the volume increases when pressure is lowered. It's important to note that Boyle's law is true only if the temperature of the gas does not change and no additional gas is added to the container or leaks out of the container.

Boyle's law states that *the volume and pressure of a sample of gas are inversely proportional to each other at constant temperature*. This statement can be expressed as follows.

Proportionality symbol. It means "is proportional to."

Proportionality constant

$$V \propto \frac{1}{P} \quad \text{and} \quad PV = k \quad \text{or} \quad V = k \frac{1}{P}$$

According to Boyle's law, when the pressure on a gas is *increased*, the volume of the gas *decreases*. For example, if the pressure is doubled, the volume decreases by half. If the volume quadruples, the pressure decreases to one-fourth of its original value.

Chemfile Mini Guide To Gas Laws

Michael Seilmaier

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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Table of Contents Chemfile Mini Guide To Gas Laws

1. Understanding the eBook Chemfile Mini Guide To Gas Laws
 - The Rise of Digital Reading Chemfile Mini Guide To Gas Laws
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemfile Mini Guide To Gas Laws
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an eBook Chemfile Mini Guide To Gas Laws
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemfile Mini Guide To Gas Laws
 - Personalized Recommendations
 - eBook Chemfile Mini Guide To Gas Laws User Reviews and Ratings

- Chemfile Mini Guide To Gas Laws and Bestseller Lists
- 5. Accessing Chemfile Mini Guide To Gas Laws Free and Paid eBooks
 - Chemfile Mini Guide To Gas Laws Public Domain eBooks
 - Chemfile Mini Guide To Gas Laws eBook Subscription Services
 - Chemfile Mini Guide To Gas Laws Budget-Friendly Options
- 6. Navigating Chemfile Mini Guide To Gas Laws eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemfile Mini Guide To Gas Laws Compatibility with Devices
 - Chemfile Mini Guide To Gas Laws Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemfile Mini Guide To Gas Laws
 - Highlighting and Note-Taking Chemfile Mini Guide To Gas Laws
 - Interactive Elements Chemfile Mini Guide To Gas Laws
- 8. Staying Engaged with Chemfile Mini Guide To Gas Laws
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemfile Mini Guide To Gas Laws
- 9. Balancing eBooks and Physical Books Chemfile Mini Guide To Gas Laws
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemfile Mini Guide To Gas Laws
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemfile Mini Guide To Gas Laws
 - Setting Reading Goals Chemfile Mini Guide To Gas Laws
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemfile Mini Guide To Gas Laws
 - Fact-Checking eBook Content of Chemfile Mini Guide To Gas Laws
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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