

# Chemistry Review Questions

## Integrating Content, Inquiry and Reasoning

Review the following data from a calorimetry experiment to determine the heat of fusion of ice. After shaking off any excess water, several ice cubes were added to 99 g of warm water contained in a calorimeter. The initial temperature of the warm water was 46.8 °C. The ice-water mixture was stirred until the temperature reached a stable, minimum value, which was 1.1 °C. Any unmelted ice remaining at this point was immediately and carefully removed using tongs and the mass of the water in the calorimeter was measured—154 g.

known  
 most warm H<sub>2</sub>O → 99g  
 T<sub>i</sub> warm H<sub>2</sub>O → 46.8 °C  
 mass mixed → 154g - 99g = 55g  
 T<sub>fin</sub> (ice) → 0 °C

T<sub>f</sub> = 1.1 °C

1. Use the heat energy equation to calculate the amount of heat in joules released by the warm water as it cooled.

$$= mC_p(\Delta T) \\
= 99 \times 4.184 \text{ J/g}^\circ\text{C} \times 45.7 \\
= 18929.7 \text{ J or } 18.93 \text{ kJ}$$

2. Based on the law of conservation of energy, what amount of heat was absorbed by the ice as it melted?

Heat absorbed → 18.93 kJ

3. Determine the amount of energy absorbed per gram of ice as it melted.

4. Calculate the heat of fusion (the heat required to melt ice) in units of kilojoules/mole.

5. The literature value for the heat of fusion of ice is 6.02 kJ/mole. What is the percent error for the experimentally determined heat of fusion?
6. When a mixture of ice and water originally at 0 °C is heated, the temperature remains constant (within experimental error) until all of the ice melts. Explain what happens to the heat energy that is absorbed during this time while the temperature does not change.

# Chemfax Ap Chemistry Laboratory 8 Answers

**David W. Hostage, Martin Fossett**



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## **Table of Contents Chemfax Ap Chemistry Laboratory 8 Answers**

1. Understanding the eBook Chemfax Ap Chemistry Laboratory 8 Answers
  - The Rise of Digital Reading Chemfax Ap Chemistry Laboratory 8 Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemfax Ap Chemistry Laboratory 8 Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemfax Ap Chemistry Laboratory 8 Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemfax Ap Chemistry Laboratory 8 Answers
  - Personalized Recommendations
  - Chemfax Ap Chemistry Laboratory 8 Answers User Reviews and Ratings

- Chemfax Ap Chemistry Laboratory 8 Answers and Bestseller Lists
- 5. Accessing Chemfax Ap Chemistry Laboratory 8 Answers Free and Paid eBooks
  - Chemfax Ap Chemistry Laboratory 8 Answers Public Domain eBooks
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  - Chemfax Ap Chemistry Laboratory 8 Answers Budget-Friendly Options
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  - Chemfax Ap Chemistry Laboratory 8 Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemfax Ap Chemistry Laboratory 8 Answers
  - Highlighting and Note-Taking Chemfax Ap Chemistry Laboratory 8 Answers
  - Interactive Elements Chemfax Ap Chemistry Laboratory 8 Answers
- 8. Staying Engaged with Chemfax Ap Chemistry Laboratory 8 Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemfax Ap Chemistry Laboratory 8 Answers
- 9. Balancing eBooks and Physical Books Chemfax Ap Chemistry Laboratory 8 Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemfax Ap Chemistry Laboratory 8 Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Chemfax Ap Chemistry Laboratory 8 Answers
  - Setting Reading Goals Chemfax Ap Chemistry Laboratory 8 Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemfax Ap Chemistry Laboratory 8 Answers
  - Fact-Checking eBook Content of Chemfax Ap Chemistry Laboratory 8 Answers
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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