

Chemistry 33

The Ideal Gas Law

(Review for Chemistry 33)

Information:

- **Boyle's Law (1662)** – Volume of a fixed mass of gas varies inversely with pressure, if the temperature is held constant.

$$P \propto \frac{1}{V} \quad \text{if constant is } n \text{ and } T$$
 where n is the number of moles of gas.
- **Charles' Law (1828)** – The volume of a gas varies directly with temperature, if the pressure is held constant.

$$V \propto T \quad \text{if constant is } n \text{ and } P$$
- **Avogadro's Hypothesis (1811)** – Samples of different gases, which contain the same number of molecules – of any complexity, size, or shape – occupy the same volume at the same temperature and pressure.

$$V \propto n \quad \text{if constant is } T \text{ and } P$$

Model 1: The Ideal Gas Law Equation.

$$PV = nRT$$

where R is a constant called the ideal gas constant.

The numerical value of the ideal gas constant is calculated from the fact that one mole of gas occupies 22.414 L at a pressure of one atmosphere and a temperature of 0°C (273.15 K):

$$R = \frac{PV}{nT} = \frac{(1 \text{ atm})(22.414 \text{ L})}{(1 \text{ mole})(273.15 \text{ K})} = \text{a value } \frac{1 \text{ atm}}{12.125}$$

Chemactivity 33 The Ideal Gas Law

SJ Ball



Chemactivity 33 The Ideal Gas Law:

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Table of Contents Chemactivity 33 The Ideal Gas Law

1. Understanding the eBook Chemactivity 33 The Ideal Gas Law
 - The Rise of Digital Reading Chemactivity 33 The Ideal Gas Law
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemactivity 33 The Ideal Gas Law
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemactivity 33 The Ideal Gas Law
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemactivity 33 The Ideal Gas Law
 - Personalized Recommendations
 - Chemactivity 33 The Ideal Gas Law User Reviews and Ratings
 - Chemactivity 33 The Ideal Gas Law and Bestseller Lists
5. Accessing Chemactivity 33 The Ideal Gas Law Free and Paid eBooks
 - Chemactivity 33 The Ideal Gas Law Public Domain eBooks
 - Chemactivity 33 The Ideal Gas Law eBook Subscription Services
 - Chemactivity 33 The Ideal Gas Law Budget-Friendly Options
6. Navigating Chemactivity 33 The Ideal Gas Law eBook Formats

- ePub, PDF, MOBI, and More
- Chemactivity 33 The Ideal Gas Law Compatibility with Devices
- Chemactivity 33 The Ideal Gas Law Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemactivity 33 The Ideal Gas Law
 - Highlighting and Note-Taking Chemactivity 33 The Ideal Gas Law
 - Interactive Elements Chemactivity 33 The Ideal Gas Law
- 8. Staying Engaged with Chemactivity 33 The Ideal Gas Law
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemactivity 33 The Ideal Gas Law
- 9. Balancing eBooks and Physical Books Chemactivity 33 The Ideal Gas Law
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemactivity 33 The Ideal Gas Law
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemactivity 33 The Ideal Gas Law
 - Setting Reading Goals Chemactivity 33 The Ideal Gas Law
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemactivity 33 The Ideal Gas Law
 - Fact-Checking eBook Content of Chemactivity 33 The Ideal Gas Law
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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