

Oxidation and Reduction Reactions: Gasoline Engines vs. Fuel Cells

(How do we recognize oxidation and reduction reactions?)

This activity requires that students know how to determine oxidation numbers.

Model 1: Examples of oxidation and reduction.

When a piece of zinc metal is placed in an aqueous solution of copper nitrate, solid copper appears deposited on the zinc metal and some of the zinc is found as $\text{Zn}^{2+}(\text{aq})$ in the solution.



Combustion of methane, a component of natural gas, with oxygen produces carbon dioxide, water and heat.




When hydrogen and oxygen gases are reacted in a hydrogen fuel cell, electrical energy is produced. The same reaction can occur when a spark initiates an explosion of hydrogen and oxygen. In either case the reaction of hydrogen and oxygen produce water.



Iron (III) oxide and aluminum react violently to produce liquid iron in this reaction.



Critical Thinking Questions

- Write the oxidation number above each atom's symbol in the four reactions (1) to (4).
- If an atom changes its oxidation number in the reaction, draw an arrow  from its symbol on the left side of the reaction to the corresponding symbol on the right side. Do this for all four reactions.
- Find and circle the reactant containing the atom whose oxidation number increases in each forward reaction.
- Find and put a square around the reactant containing the atom whose oxidation number decreases in each forward reaction.

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Organic Chemistry, a Guided Inquiry Andrei Straumanis, 2004

Chemical Abstracts, 1907

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