

4. A student adds 15.0 mL of water to a 25.0 mL sample of 0.80 M Li_3PO_4 solution.

$$FC = IC \times \frac{V}{FV}$$

- a) Calculate the final $[\text{Li}^+]$ (2 marks)

$$\text{Final } [\text{Li}_3\text{PO}_4] = 0.80 \text{ M} \times \frac{25.0 \text{ mL}}{40.0 \text{ mL}} = 0.50 \text{ M}$$



$$\text{final } [\text{Li}^+] = 1.5 \text{ M}$$

- b) Calculate the final $[\text{PO}_4^{3-}]$ (1 mark)



$$\text{final } [\text{PO}_4^{3-}] = 0.50 \text{ M}$$

5. Calculate the mass of solid K_2CO_3 which must be added to 500.0 mL of water in order to make a solution in which the $[\text{K}^+]$ is 3.0 M. (3 marks)



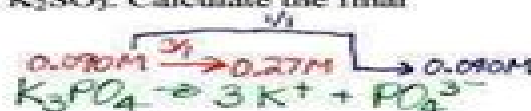
$$\text{mol} = \text{M} \times \text{L} = 1.50 \text{ M} \times 0.5000 \text{ L} = 0.750 \text{ mol } \text{K}_2\text{CO}_3$$

$$\text{mass} : 0.750 \text{ mol } \text{K}_2\text{CO}_3 \times \frac{138.2 \text{ g } \text{K}_2\text{CO}_3}{1 \text{ mol } \text{K}_2\text{CO}_3} = 103.65 \text{ g}$$

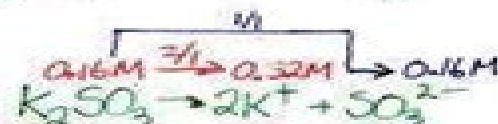
$$\text{mass } \text{K}_2\text{CO}_3 = 104 \text{ g}$$

6. 60.0 mL of 0.45 M K_3PO_4 is mixed with 240.0 mL of 0.20 M K_2SO_3 . Calculate the final concentration of all three ions in the solution. (3 marks)

$$\text{final } [\text{K}_3\text{PO}_4] = 0.45 \text{ M} \times \frac{60.0 \text{ mL}}{300.0 \text{ mL}} = 0.090 \text{ M}$$



$$\text{final } [\text{K}_2\text{SO}_3] = 0.20 \text{ M} \times \frac{240.0 \text{ mL}}{300.0 \text{ mL}} = 0.16 \text{ M}$$



$$\text{final } [\text{K}^+] = 0.27 \text{ M} + 0.32 \text{ M} = 0.59 \text{ M}$$

$$\text{Answer: } [\text{PO}_4^{3-}] = 0.090 \text{ M} \quad [\text{SO}_3^{2-}] = 0.16 \text{ M} \quad [\text{K}^+] = 0.59 \text{ M}$$

Chem Fax Acid Base Titrations Answers

Clifford Lane



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- Chapter 3: Chem Fax Acid Base Titrations Answers in Everyday Life
- Chapter 4: Chem Fax Acid Base Titrations Answers in Specific Contexts
- Chapter 5: Conclusion

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Table of Contents Chem Fax Acid Base Titrations Answers

1. Understanding the eBook Chem Fax Acid Base Titrations Answers
 - The Rise of Digital Reading Chem Fax Acid Base Titrations Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chem Fax Acid Base Titrations Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an eBook Platform
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chem Fax Acid Base Titrations Answers
 - Personalized Recommendations
 - eBook Reviews and Ratings
 - Bestseller Lists
5. Accessing Chem Fax Acid Base Titrations Answers Free and Paid eBooks
 - Public Domain eBooks
 - eBook Subscription Services
 - Budget-Friendly Options
6. Navigating eBook Formats
 - ePub, PDF, MOBI, and More
 - eBook Compatibility with Devices
 - Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes
 - Highlighting and Note-Taking
 - Interactive Elements

8. Staying Engaged with Chem Fax Acid Base Titrations Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chem Fax Acid Base Titrations Answers
9. Balancing eBooks and Physical Books Chem Fax Acid Base Titrations Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chem Fax Acid Base Titrations Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chem Fax Acid Base Titrations Answers
 - Setting Reading Goals Chem Fax Acid Base Titrations Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chem Fax Acid Base Titrations Answers
 - Fact-Checking eBook Content of Chem Fax Acid Base Titrations Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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