

**Ideal Gas Law,  $PV = nRT$** 

Name \_\_\_\_\_

Instructor \_\_\_\_\_ Section \_\_\_\_\_ Partner \_\_\_\_\_

(Unless otherwise noted, 2 pts. for each blank)

|                                                                       | Data, Trial 1                                                  | Data, Trial 2     |
|-----------------------------------------------------------------------|----------------------------------------------------------------|-------------------|
| Molarity of HCl (M) (1pt. each):                                      | _____                                                          | _____             |
| Volume of HCl (mL):                                                   | <u>19.9 mL</u>                                                 | <u>20.0 mL</u>    |
| Mass of Mg (gram):                                                    | <u>0.044 g</u>                                                 | <u>0.046 g</u>    |
| *Volume of gas before placing in<br>Equalization Chamber (mL) (1 pt.) | <u>45.1 mL</u><br>(*This value NOT used in calculations below) | <u>46.1 mL</u>    |
| Volume of gas after placing<br>in Equalization Chamber (mL):          | <u>44.9 mL</u>                                                 | <u>46.0 mL</u>    |
| Barometric Pressure (mm Hg):                                          | <u>738.2 mmHg</u>                                              | <u>738.2 mmHg</u> |
| Temperature (°C):                                                     | <u>22.0 °C</u>                                                 | <u>22.0 °C</u>    |

Write the balanced overall equation for the reaction of magnesium and HCl to form hydrogen gas and the magnesium salt (2 pts):

|                                                                                          |       |       |
|------------------------------------------------------------------------------------------|-------|-------|
| Mole of Mg reacted                                                                       | _____ | _____ |
| Mole of H <sub>2</sub> formed (based on the<br>stoichiometric equation for the reaction) | _____ | _____ |
| Vapor Pressure of water from curve in mmHg:                                              | _____ | _____ |
| Pressure of H <sub>2</sub> from Dalton's Law in mmHg:                                    | _____ | _____ |
| Pressure of H <sub>2</sub> in atm                                                        | _____ | _____ |
| Volume of H <sub>2</sub> in liter<br>(Using volume measured in Equalization Chamber)     | _____ | _____ |
| Temperature in Kelvin                                                                    | _____ | _____ |
| Calculation of R (6 pts. each)                                                           | _____ | _____ |
| Percent error                                                                            | _____ | _____ |
| Average R =                                                                              | _____ | _____ |

# Chem 1211 Lab Manual Answers

**VM Jensen**



## **Chem 1211 Lab Manual Answers:**

## Decoding **Chem 1211 Lab Manual Answers**: Revealing the Captivating Potential of Verbal Expression

In an era characterized by interconnectedness and an insatiable thirst for knowledge, the captivating potential of verbal expression has emerged as a formidable force. Its power to evoke sentiments, stimulate introspection, and incite profound transformations is genuinely awe-inspiring. Within the pages of "**Chem 1211 Lab Manual Answers**," a mesmerizing literary creation penned by a celebrated wordsmith, readers attempt an enlightening odyssey, unraveling the intricate significance of language and its enduring effect on our lives. In this appraisal, we shall explore the book's central themes, evaluate its distinctive writing style, and gauge its pervasive influence on the hearts and minds of its readership.

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