

11 Which compound has the greatest percent composition by mass of sulfur?

- (1) BaS (2) CaS (3) MgS (4) SrS

way #64

12 Two molecules of HBr collide and then form H_2 and Br_2 . During the collision, the bonds in the HBr molecules are

- (1) broken as energy is absorbed
(2) broken as energy is released
(3) formed as energy is absorbed
(4) formed as energy is released

way #43

BARE

13 Which atom in the ground state has a stable electron configuration?

- (1) carbon (2) magnesium (3) neon (4) oxygen

way #49

14 Which statement describes a multiple covalent bond?

- (1) Two electrons are shared.
(2) Four electrons are shared.
(3) Two electrons are transferred.
(4) Four electrons are transferred.

way #46

15 The electronegativity difference between the atoms in a molecule of HCl can be used to determine

- (1) the entropy of the atoms
(2) the atomic number of the atoms
(3) the first ionization energy of the atoms
(4) the polarity of the bond between the two atoms

way #51

16 Which two gases can not be broken down by chemical means?

- (1) CO and He (2) CO and NH_3 (3) Xe and He (4) Xe and NH_3

way #5

elements Table S

17 Two substances in a mixture differ in density and particle size. These properties can be used to

- (1) separate the substances
(2) chemically combine the substances
(3) determine the freezing point of the mixture
(4) predict the electrical conductivity of the mixture

18 Which unit is used to express an amount of thermal energy?

- (1) gram (2) mole (3) joule (4) pascal

Table D

19 Under which conditions of temperature and pressure does a real gas behave most like an ideal gas?

- (1) low temperature and low pressure
(2) low temperature and high pressure
(3) high temperature and low pressure
(4) high temperature and high pressure

way #9

20 According to the kinetic molecular theory for an ideal gas, all gas particles

- (1) are in random, constant, straight-line motion
(2) are separated by very small distances relative to their sizes
(3) have strong intermolecular forces
(4) have collisions that decrease the total energy of the system

way #8

21 Which mathematical expression represents the heat of reaction for a chemical reaction?

- (1) (the heat of fusion) - (the heat of vaporization)
(2) (the heat of vaporization) - (the heat of fusion)
(3) (the potential energy of the products) - (the potential energy of the reactants)
(4) (the potential energy of the reactants) - (the potential energy of the products)

way #78

22 At 101.3 kPa and 298 K, a 1.0-mole sample of which compound absorbs the greatest amount of heat as the entire sample dissolves in water?

- (1) LiBr -48.83 (2) NaCl +2.38 (3) NaOH -46.51 (4) NH_4Cl +14.8

way #11

Table I

23 For a reaction at equilibrium, which change can increase the rates of the forward and reverse reactions?

- (1) a decrease in the concentration of the reactants
(2) a decrease in the surface area of the products
(3) an increase in the temperature of the system
(4) an increase in the activation energy of the forward reaction

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S. Chand Experts

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Analysis, Synthesis, and Design of Chemical Processes Richard Turton, Joseph A. Shaeiwitz, Debangsu Bhattacharyya, Wallace B. Whiting, 2018-06-15 The Leading Integrated Chemical Process Design Guide With Extensive Coverage of Equipment Design and Other Key Topics More than ever effective design is the focal point of sound chemical engineering Analysis Synthesis and Design of Chemical Processes Fifth Edition presents design as a creative process that integrates the big picture and small details and knows which to stress when and why Realistic from start to finish it moves readers beyond classroom exercises into open ended real world problem solving The authors introduce up to date integrated techniques ranging from finance to operations and new plant design to existing process optimization The fifth edition includes updated safety and ethics resources and economic factors indices as well as an extensive new section focused on process equipment design and performance covering equipment design for common unit operations such as fluid flow heat transfer separations reactors and more Conceptualization and analysis process diagrams configurations batch processing product design and analyzing existing processes Economic analysis estimating fixed capital investment and manufacturing costs measuring process profitability and more Synthesis and optimization process simulation thermodynamic models separation operations heat integration steady state and dynamic process simulators and process regulation Chemical equipment design and performance a full section of expanded and revamped coverage of designing process equipment and evaluating the performance of current equipment Advanced steady state simulation goals models solution strategies and sensitivity and

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Sager,Takashi Sano,Elisabetta Erba,2015-05 The origin evolution and environmental impact of large igneous provinces LIPs represents a topic of high scientific importance because the magmatism associated with these features cannot be directly related to plate tectonics and because the eruption of flood basalts may have global environmental consequences Oceanic LIPs are even more poorly understood due to their relative inaccessibility This volume takes a multidisciplinary approach to understanding LIP origin evolution and environmental impact in ocean basins Papers that focus on plate tectonic reconstructions petrologic and geophysical investigations of various LIPs and sedimentological and micropaleontological evidence of syn LIP sediments are presented Precious materials and data from dredging cruises and scientific ocean drilling expeditions have made this volume possible **El-Hi Textbooks in Print** ,1984 **Sales Management** Thomas N.

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