

# CHAPTER 17 STUDY GUIDE

## Chemical Equilibrium

### Section 17.1 Equilibrium: A State of Dynamic Balance

*In your textbook, read about chemical equilibrium.*

**Complete each statement.**

1. When a reaction results in almost complete conversion of reactants to products, chemists say the reaction goes to \_\_\_\_\_.
2. A reaction that can occur in both the forward and the reverse directions is called a(n) \_\_\_\_\_.
3. \_\_\_\_\_ is a state in which the forward and reverse reactions balance each other because they take place at equal rates.
4. At equilibrium, the concentrations of reactants and products are \_\_\_\_\_ but that does not mean that the amounts or concentrations are \_\_\_\_\_.
5. Equilibrium is a state of \_\_\_\_\_, not one of \_\_\_\_\_.

*In your textbook, read about equilibrium expressions and constants.*

**For each statement below, write *true* or *false*.**

- \_\_\_\_\_ 6. The law of chemical equilibrium states that at a given pressure, a chemical system may reach a state in which a particular ratio of reactant to product concentrations has a constant value.
- \_\_\_\_\_ 7. The equation  $\text{H}_2(\text{g}) + \text{I}_2(\text{g}) \rightleftharpoons 2\text{HI}(\text{g})$  is an example of a homogeneous equilibrium.
- \_\_\_\_\_ 8. If an equilibrium constant has a value less than one, the reactants are favored at equilibrium.
- \_\_\_\_\_ 9. The value for  $K_{\text{eq}}$  is constant only at a specific volume.
- \_\_\_\_\_ 10. If the equilibrium constant for a reaction at 300 K is 49.7, the concentration of the reactants will be greater than the concentration of the products.
- \_\_\_\_\_ 11. A heterogeneous equilibrium means that reactants and products are present in more than one state.
- \_\_\_\_\_ 12. The product of the forward chemical reaction is HI, for the equilibrium expression:

$$K_{\text{eq}} = \frac{[\text{HI}]^2}{[\text{H}_2][\text{I}_2]}$$

# Chemistry Study Guide Answers Texas

**Chao Zhang**



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*Study Guide and Solutions Manual for Seager/Slabaugh's Chemistry for Today* Seager/Slabaugh, 2004 The fifth edition of the Study Guide and Student Solutions Manual has been updated to reflect all of the changes to the text This ancillary tests the student on the learning objectives in each chapter and provides answers to all of the even numbered end of chapter exercises New additional activities have been added to include a review of each section of the chapter and a section entitled Tying It All Together with a Laboratory Application *Student's Guide to Fundamentals of Chemistry* Jo A.

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