BASIC CHEMISTRY

CHEMISTRY

- ✓ Medieval Latin "Alchimia"/ Arabic "Alkimya", the chemistry
- ✓ The study of the composition and structure of the changes in matter
- It is often called the central science because knowledge in this field is a prerequisite for all the other physical and biological sciences.

HISTORICAL DEVELOPMENT OF THE ATOMIC STRUCTURE

ATOM

- It is the smallest possible unit into which matter can be divided, while still maintaining its properties
- ✓ Building blocks of matter
- ✓ It has protons (+), neutrons (0), and electons (-)
 - Democritus was a greek philosopher who first introduced the concept of atoms. He believed that matter is comprises of tiny indivisible particles known as atomos.
 - 2. Dalton's Atomic Theory
 - a. All matter is made up of indivisible particles known as atoms.
 - b. All atoms of the same elements are alike.
 - c. Compounds are formed when atoms of different elements combine in certain proportions.
 - d. atoms are only rearranged but neither created nor destroyed in chemical reactions.
 - Law of Definite Proportion a given compound always contains the same elements in exactly the same proportion by mass.
 - H₂0 ratio of two hydrogen atom with one oxygen atom only. If it becomes ratio of two hydrogen atom with two oxygen atom, it is not already water.
 - 4. Law of Multiple Proportions when two elements can combine to form more than one type of compound, the masses of one element that combine with a fixed mass of the other elements are in a ratio of small whole numbers.

CaO + CO₂ → CaCO₃

- Law of Conservation of Mass states that matter can neither be created nor destroyed. Matter is just transformed.
- 6. The discovery of the different sub-atomic particles:
 - J.J. Thompson determined the charge to mass ratio of the electron using a cathode ray tube.

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CO Houle

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Study Guide to Accompany Basics for Chemistry Martha Mackin, 2012-12-02 Study Guide to Accompany Basics for Chemistry is an 18 chapter text designed to be used with Basics for Chemistry textbook Each chapter contains Overview Topical Outline Skills and Common Mistakes which are all keyed to the textbook for easy cross reference The Overview section summarizes the content of the chapter and includes a comprehensive listing of terms a summary of general concepts and a list of numerical exercises while the Topical Outline provides the subtopic heads that carry the corresponding chapter and section numbers as they appear in the textbook The Fill in Multiple Choice are two sets of questions that include every concept and numerical exercise introduced in the chapter and the Skills section provides developed exercises to apply the new concepts in the chapter to particular examples The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students Organic Chemistry Study Guide with Solutions Manual Neil E. Schore, 2007 The guide includes chapter introductions that highlight new material chapter outlines detailed comments for each chapter section a glossary and solutions to the end of chapter problems presented in a way that shows students how to reason their way to the answer

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