

MCQ FOR BOARD EXAMINATION

XI Chapter – 01

STOICHIOMETRY

- If the volume occupied by oxygen gas (O_2) at STP is 44.8 dm^3 , the number of molecules of O_2 in the Vessels are:
 (a) 12.01×10^{24} (b) 12.01×10^{25} (c) 12.01×10^{23} (d) 12.01×10^{22}
- The number of Carbon atoms in 1 mole of sugar ($C_{12}H_{22}O_{11}$) are approximately:
 (a) 12.01×10^{23} (b) 12.01×10^{24} (c) 12.01×10^{25} (d) 72×10^{23}
- In the reaction $2Na + 2H_2O \rightarrow 2NaOH + H_2$, if 23g of Na reacts with excess of water, the volume of hydrogen gas (H_2) liberated at STP should be:
 (a) 11.2 dm^3 (b) 2.44 dm^3 (c) 10.60 dm^3 (d) 44.8 dm^3
- Which of the following sample of substance contains the same number of atoms as that of 20g calcium:
 (a) 16g S (b) 32g O (c) 39g K (d) 48g Mg
- Which of the following statement is incorrect?
 (a) The mass of 1 mole Cl_2 gas is 35.5g
 (b) One mole of H_2 gas contains 6.01×10^{23} molecules of H_2
 (c) Number of atoms in 23g Na and 24g of Mg are equal
 (d) One moles of O_2 at S.T.P occupy 22.4 dm^3
- For Avogadro's number, this statement is incorrect:
 (a) It is the number of particles in one mole of any substances
 (b) Its numerical value is 6.01×10^{23}
 (c) Its value change if temperature increases
 (d) Its value change if number of moles increases
- The minimum number of moles are present in:
 (a) 1 dm^3 of methane gas at STP (b) 5 dm^3 of helium gas at STP
 (c) 10 dm^3 of hydrogen gas at STP (d) 22.4 dm^3 of chlorine gas at STP
- Number of atoms in 60g of Carbon are:
 (a) 12.01×10^{23} (b) 3.01×10^{24} (c) 6.01×10^{25} (d) 6.01×10^{24}
- Maximum number of molecules present in the following sample of gas:
 (a) 100g O_2 (b) 100g of CH_4
 (c) 100g CO_2 (d) 100g of Cl_2
- Generally actual yield is:
 (a) Greater than theoretical yield (b) Less than theoretical yield
 (c) Equal to the theoretical yield (d) Some times greater and sometimes less than theoretical yield
- The short hand disruption of a very large or small number in the power of 10 is called:
 (a) Exponential notation (b) Rounding off (c) Significant figure (d) All of them
- 60000 is equal to:
 (a) 6×10^5 (b) 6.00×10^4 (c) 60×10^4 (d) 6.0×10^3
- 602000 in exponential notation may be expressed as:
 (a) 6.02×10^5 (b) 6.02×10^{-5} (c) 0.602×10^5 (d) 6.02×10^{-4}
- 330000 in exponential notation may be expressed as:
 (a) 3.3×10^6 (b) 0.33×10^5 (c) $3.3 \times 10^{+5}$ (d) 3.3×10^{-5}
- The number 683×10^{-7} expressed in scientific notation as
 (a) 68.3×10^{-7} (b) 6.83×10^{-5} (c) 6.83×10^{-11} (d) 6.83×10^{-12}
- The reactant which does not completely consume in the reaction and some of its amount remains unused at the end of reaction.
 (a) Reactant in excess (b) Limiting reactant (c) Product in excess (d) Reaction
- 1 mole of Gold (Au) 197g how many no of atom.
 (a) 6.02×10^{23} atom (b) 12.02×10^{23} atom (c) 6.02×10^{25} atom (d) 12.02×10^{25} atom

Chemistry Stoichiometry Mcqs

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