

CHEMISTRY TERMINOLOGY

Term	Definition
abrasive	A very hard, brittle, heat-resistant substance that is used to grind the edges or rough surfaces of an object. boron carbide, diamond, and corundum are abrasives.
absolute error	The uncertainty in a measurement, expressed with appropriate units. For example, if three replicate weights for an object are 1.00 g, 1.05 g, and 0.95 g, the absolute error can be expressed as ± 0.05 g. Absolute error is also used to express inaccuracies; for example, if the "true value" is 1.11 g and the measured value is 1.00 g, the absolute error could be written as $1.00 \text{ g} - 1.11 \text{ g} = -0.11 \text{ g}$. Note that when absolute errors are associated with indeterminate errors, they are preceded with "±"; when they are associated with determinate errors, they are preceded by their sign.
absolute zero	The temperature at which the volume of an ideal gas becomes zero; a theoretical coldest temperature that can be approached but never reached. Absolute zero is zero on the Kelvin scale, -273.15°C on the Celsius scale, and -459.67°F on the Fahrenheit scale.
absorbance	A measure of the amount of light absorbed by a sample. The absorbance (A) equals minus the base-10 log of the transmittance.
absorption	1. Penetration of molecules into the bulk of a solid or liquid, forming either a solution or compound. Absorption can be a chemical process (a strong solution of NaOH absorbs CO_2 from the air) or a physical process (palladium absorbs hydrogen gas). 2. Capture and transformation of energy by a substance; for example, copper looks reddish because it absorbs blue light. An absorbent captures another material and distributes it throughout; an adsorbent captures another material and distributes it on its surface only.
absorptivity	The absorbance of a solution per unit of path length and per unit concentration; $a = A/(bc)$ where a, A, b, and c are the absorptivity, absorbance, path length, and concentration, respectively. Absorptivity varies with wavelength of the incident light.
accelerator	1. A substance that makes vulcanization of rubber occur more quickly or at a lower temperature. 2. A substance that makes crosslinking in a polymer occur more quickly or at a lower temperature, e. g., accelerators are added to Super Glue to make it set up quickly.
accuracy	Accuracy is the correctness of a single measurement. The accuracy of a measurement is assessed by comparing the measurement with the true or accepted value, based on evidence independent of the measurement. The closeness of an average to a true value is referred to as "trueness".
acid	Nonmetallic oxides or organic compounds that react with water to form acids*. For example, SO_2 , CO_2 , P_2O_5 , and SO_3 are the acid anhydrides of sulfurous, carbonic, phosphoric, and sulfuric acids, respectively. Acetic anhydride ($(\text{CH}_3\text{CO})_2\text{O}$) reacts with water to form acetic acid.
acidic solution	A solution in which the hydrogen ion* activity* is higher than that of the hydroxide ion*, when the solvent* is water.
actinide	Elements 89-102 are called actinides. Electrons added during the Aufbau construction* of actinide atoms go into the 5f subshell. Actinides are unstable and undergo radioactive decay*. The most common actinides on Earth are uranium and thorium.
adhesion	Attraction between different substances on either side of a phase boundary.
adsorbent	A substance that collects molecules of another substance on its surface. For example, gases that make water taste bad are strongly adsorbed on activated charcoal* granules in water filters.
aerosol	A colloid* in which solid particles or liquid droplets are suspended in a gas. Smoke is an example of a solid aerosol; fog is an example of a liquid aerosol.
alkaline	Having a pH* greater than 7
amino acid	Amino acids are molecules that contain at least one amine group ($-\text{NH}_2$) and at least one carboxylic acid group ($-\text{COOH}$). When these groups are both attached to the same carbon, the acid is an alpha-amino acid. alpha-amino acids are the basic building blocks of proteins.

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