

**THE FOLLOWING QUESTIONS ARE PRACTICE QUESTIONS FOR THE CSN CHEMISTRY PLACEMENT EXAM**

The chemistry placement test is used to assess your preparation for General Chemistry I, CHEM 121 at CSN. The test is a standardized test that will examine your knowledge of chemistry and basic mathematical skills. You will be provided scratch paper, a Periodic Table of the Elements and a nonprogrammable scientific calculator. The following sample questions are intended only as a guide to show you the relative level of difficulty and example topics covered in the Chemistry Placement Exam. The actual exam will have a time limit of 45 minutes to answer 44 multiple choice questions. Click the following link to view a [Periodic Table](#).

**MULTIPLE CHOICE.** Choose the single choice that best answers the question.

- 1) How many total atoms are in the formula



A) 12                      B) 14                      C) 9                      D) 8

- 2) How many electrons are in  $\text{Br}^-$ ?

A) 4                      B) 7                      C) 36                      D) 34

- 3) What are the coefficients for the following reaction when it is properly balanced?



A) 2, 3, 6, 1                      B) 6, 1, 3, 2  
C) 2, 3, 1, 6                      D) 2, 1, 1, 3

- 4) Given that  $4 \text{ NH}_3 + 5 \text{ O}_2 \rightarrow 4 \text{ NO} + 6 \text{ H}_2\text{O}$ , when 4.50 mol of  $\text{H}_2\text{O}$  are formed, the amount of NO formed is

A) 3.00 mol                      B) 6.75 mol  
C) 1.50 mol                      D) 4.50 mol

- 5) What type of a reaction occurs when a sodium hydroxide solution is mixed with an acetic acid solution?

A) gas evolution  
B) oxidation-reduction  
C) acid-base neutralization  
D) precipitation

- 6) Which of the following atoms has the smallest radius?

A) Ne                      B) O  
C) Li                      D) B

- 7) What is the molecular geometry if you have a double bond, a single bond and 1 lone pair around the central atom?

A) trigonal pyramidal  
B) linear  
C) tetrahedral  
D) bent

- 8) What volume of 12.0 M HCl is required to make 75.0 mL of 3.50 M HCl?

A) 0.560 mL                      B) 257 mL  
C) 560. mL                      D) 21.9 mL

$$9) \frac{(2.80 \times 10^{-3})(1.28 \times 10^1)}{(3.20 \times 10^{-4})} = \underline{\hspace{2cm}}$$

A)  $1.27 \times 10^1$                       B)  $2.0 \times 10^2$   
C)  $1.2 \times 10^6$                       D)  $1.12 \times 10^2$

- 10) Of the molecules below, only \_\_\_\_\_ is polar.

A)  $\text{CH}_4$                       B)  $\text{SbF}_5$   
C)  $\text{CO}_2$                       D)  $\text{AsH}_3$

- 11) Which one of the following is a metalloid?

A) S                      B) Ge                      C) C                      D) Pb

- 12) A certain atom has a charge of -2 and has 16 protons and 16 neutrons in the nucleus. How many electrons does this atom possess?

A) 34                      B) 14                      C) 32                      D) 18

- 13) How many moles of pyridine ( $\text{C}_5\text{H}_5\text{N}$ ) are contained in 3.13 g of pyridine?

A) 0.00404                      B) 25.3  
C) 0.0396                      D) 0.319

- 14) Solve the following equation for x:

$$4x - b = 2x + c$$

A)  $c - b - 6$                       B)  $\frac{c + b}{2}$   
C)  $\frac{c + b}{6}$                       D)  $\frac{c - b}{2}$

# Chemistry Placement Test Study Guide Texas Tech

**Clemens Wendtner**



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*Energy-related Environment and Safety Research for ...* ,1980      **Inventory of Federal Energy-related Environment and Safety Research for FY 1979** ,1980      **Understanding Risk to Wildlife from Exposures to Per- and Polyfluorinated Alkyl Substances (PFAS)** Mark S. Johnson,Michael J. Quinn Jr.,Marc A. Williams,Allison M. Narizzano,2021-04-22

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