

## Unit 2 Vocabulary - The Chemistry of Life

<b>Nucleus</b>	The positively charged core of hydrogen ( $H^+$ ), found in acids.
<b>Covalent Bond</b>	Negatively charged molecule of one oxygen bonded to one hydrogen ( $OH^-$ ), found in bases.
<b>pH Scale</b>	A molecule that exists in all three states of matter (solid, liquid and gas) on Earth, is the primary fluid of all living organisms.
<b>Isotope</b>	Tells us the number of atoms of each element in a compound, i.e. $NaCl$ is one sodium bonded to one chlorine and forming sodium chloride or salt.
<b>Ionic Bond</b>	The transfer of valence electrons between atoms, formed between a metal and a nonmetal, the strongest (most polar) of all bond types.
<b>Base</b>	A process that involves the changing from one substance into something completely new or different.
<b>Ion</b>	A particle smaller than an atom, protons, neutrons and electrons.
<b>Atoms</b>	When electrons are bound to each other forming a compound (ionic or covalent) are the most common type.
<b>Chemical Reaction</b>	The amount of a substance in a specified space.
<b>Acid</b>	Any substance that forms hydroxide ions ( $OH^-$ ) in water, has a pH greater than 7.
<b>Subatomic Particle</b>	The positively charged center of an atom, contains protons ( $p^+$ ) and neutrons (neutrons).
<b>Chemical Bond</b>	A measure of how acidic or basic a solution is, ranges from 0 to 14, 7 is neutral, <7 is acidic, >7 is basic.
<b>Concentration</b>	Atoms of the same element that have a different number of neutrons in the nucleus.
<b>Hydrogen Ion</b>	The smallest particle of something, the basic building blocks of all matter.
<b>Hydroxide Ion</b>	Chemical bond formed when two atoms combine by the sharing of pairs of valence electrons.
<b>Water</b>	Any atom that has a charge, created by the loss ( $p^+$ ) or gain ( $e^-$ ) of electrons.
<b>Chemical Formula</b>	Any substance that forms hydrogen ( $H^+$ ) ions in water, has a pH less than 7.

# Chemistry Of Life Vocabulary Practice Answers

**J Spring**



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