

Chemical Monitoring and Management

9.4.1 Much of the work of chemists involves monitoring the reactants and products of reactions and managing reaction conditions

Outline the role of a chemist employed in a named industry or enterprise, identifying the branch of chemistry undertaken by the chemist and explaining a chemical principle that the chemist uses. Gather, process and present information from secondary sources about the work of practising scientists identifying:

- the variety of chemical occupations
- a specific chemical occupation for a more detailed study

An analytical chemist working for Sydney Water will use AAS to monitor concentrations of metals, such as Pb and Hg, in water samples which eventually will be supplied to households. The principle of chemistry used by the scientist is that metal electrons move from one energy shell to a higher one by absorbing electromagnetic radiation of a wavelength specific to that metal, giving each metal a unique emission spectrum \therefore metals can be identified and their concentration can be determined even in the presence of other metals. It is important that the scientist monitor and regulate the concentrations of metals in drinking water sources as excessive concentrations can lead to poisoning and illness, e.g. lead poisoning can cause mental retardation, but trace element metals, like Zn, are essential in small concentrations for body functioning.

Identify the need for collaboration between chemists as they collect and analyse data

- Chemists specialising in various fields of chemistry, e.g. analytical, organic, industrial, need to collaborate to solve problems requiring a broad spectrum of chemical knowledge.
- Chemists need to regularly exchange viewpoints and have an open minded but critical approach to ensure that our scientific knowledge is constantly improving.

Describe an example of a chemical reaction such as combustion, where reactants form different products under different conditions and thus would need monitoring

Combustion reactions can produce solely carbon dioxide or a mixture of poisonous carbon monoxide and pollutant soot (C), depending on the amount of oxygen present. Using methane as an example:

Oxygen Supply	Bunsen Burner Hole	Flame Colour	Rate of Combustion	Reaction Equation
Sufficient	Open	Blue	Complete combustion; maximum energy released	$\text{CH}_4 (\text{g}) + 2\text{O}_2 (\text{g}) \rightarrow \text{CO}_2 (\text{g}) + 2\text{H}_2\text{O} (\text{l})$
Limited	Partially open	Mauve	Incomplete combustion	$2\text{CH}_4 (\text{g}) + 3\text{O}_2 (\text{g}) \rightarrow 2\text{CO} (\text{g}) + 4\text{H}_2\text{O} (\text{l})$
Blocked	Closed	Yellow	Incomplete combustion	$\text{CH}_4 (\text{g}) + \text{O}_2 (\text{g}) \rightarrow \text{C} (\text{s}) + 2\text{H}_2\text{O} (\text{l})$

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