

Competition File

- B3.** In Haber process, 30 L of dihydrogen and 30 L of dinitrogen were taken for reaction which yielded only 50% of the expected product. What will be the composition of the gaseous mixture under the aforesaid conditions in the end?
 (a) 20 L NH_3 , 25 L N_2 and 20 L H_2
 (b) 10 L NH_3 , 25 L N_2 and 15 L H_2
 (c) 20 L NH_3 , 10 L N_2 and 30 L H_2
 (d) 20 L NH_3 , 25 L N_2 and 15 L H_2
 (C.B.S.E. Med. 2003)
- B4.** 6.02×10^{20} molecules of urea are present in 100 mL of its solution. The concentration of solution is
 (a) 0.02 M (b) 0.01 M
 (c) 0.001 M (d) 0.1 M (A.I.E.E.E. 2004)
- B5.** One mole of magnesium nitride on reaction with an excess of water gives:
 (a) two moles of ammonia
 (b) one mole of nitric acid
 (c) one mole of ammonia
 (d) two moles of nitric acid (A.I.E.E.E. 2004)
- B6.** The maximum number of molecules is present in:
 (a) 5 L of N_2 gas at STP
 (b) 0.5 g of H_2 gas
 (c) 10 g of O_2 gas
 (d) 15 L of H_2 gas at STP (C.B.S.E. Med. 2004)
- B7.** A sample of phosphorus trichloride (PCl_3) contains 1.4 moles of the substance. How many atoms are there in the sample?
 (a) 4 (b) 5.6
 (c) 8.431×10^{23} (d) 3.372×10^{24}
 (e) 2.409×10^{24} (Kerala P.M.T. 2004)
- B8.** A gaseous mixture contains 50% He and 50% CH_4 by volume. What is the percent by weight of CH_4 in the mixture?
 (a) 19.97% (b) 20.05%
 (c) 50% (d) 75%
 (e) 80.03% (Kerala P.M.T. 2004)
- B9.** The vapour density of gas A is four times that of B. If molecular mass of B is M, then molecular mass of A is:
 (a) 4 M (b) 2 M
 (c) M (d) $\frac{M}{4}$ (B.H.U. 2004)
- B10.** The mass of carbon anode consumed (giving only carbon dioxide) in the production of 270 kg of aluminium metal from bauxite by Hall process is:
 (a) 180 kg (b) 270 kg
 (c) 540 kg (d) 90 kg (C.B.S.E. Med. 2005)
- B11.** If 30 mL of H_2 and 20 mL of O_2 react to form water, what is left at the end of the reaction?
 (a) 10 mL of H_2 (b) 5 mL of H_2
 (c) 10 mL of O_2 (d) 5 mL of O_2 (A.F.M.C. 2005)
- B12.** An alkaloid contains 17.28% of nitrogen and its molecular mass is 162. The number of nitrogen atoms present in one molecule of alkaloid is
 (a) five (b) four
 (c) three (d) two
 (e) one (Kerala P.M.T. 2005)
- B13.** For the formation of 3.65 g of hydrogen chloride gas, what volumes of hydrogen gas and chlorine gas are required at N.T.P. conditions?
 (a) 1.12 L, 1.12 L (b) 1.12 L, 2.4 L
 (c) 3.65 L, 1.83 L (d) 1 L, 1 L
 (e) 3.65 L, 1.83 L (Kerala C.E.T. 2005)
- B14.** If we consider that $\frac{1}{6}$ in place of $\frac{1}{12}$ mass of carbon atom is taken to be the relative atomic mass unit, the mass of one mole of a substance will
 (a) decrease twice
 (b) increase two fold
 (c) remain unchanged
 (d) be a fraction of molecular mass of the substance (A.I.E.E.E. 2005)
- B15.** How many moles of magnesium phosphate, $\text{Mg}_3(\text{PO}_4)_2$ will contain 0.25 mole of oxygen atoms?
 (a) 3.125×10^{-2} (b) 1.25×10^{-2}
 (c) 2.5×10^{-2} (d) 0.02 (A.I.E.E.E. 2006)
- B16.** The crystalline salt $\text{Na}_2\text{SO}_4 \cdot x\text{H}_2\text{O}$ on heating loses 55.9 % of its weight. The formula of crystalline salt is
 (a) $\text{Na}_2\text{SO}_4 \cdot 5\text{H}_2\text{O}$ (b) $\text{Na}_2\text{SO}_4 \cdot 7\text{H}_2\text{O}$
 (c) $\text{Na}_2\text{SO}_4 \cdot 2\text{H}_2\text{O}$ (d) $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$
 (e) $\text{Na}_2\text{SO}_4 \cdot 6\text{H}_2\text{O}$ (Kerala P.M.T. 2007)
- B17.** In the reaction:
 $2\text{Al}(s) + 6\text{HCl}(aq) \longrightarrow 2\text{Al}^{3+}(aq) + 6\text{Cl}^-(aq) + 3\text{H}_2(g)$
 (a) 33.6 L $\text{H}_2(g)$ is produced regardless of temperature and pressure for every mole of Al that reacts.
 (b) 67.2 L $\text{H}_2(g)$ at STP is produced for every mole of Al that reacts.
 (c) 11.2 L $\text{H}_2(g)$ at STP is produced for every mole of $\text{HCl}(aq)$ consumed.
 (d) 6 L $\text{HCl}(aq)$ is consumed for every 3 L of $\text{H}_2(g)$ produced. (A.I.E.E.E. 2007)
- B18.** 80 g of oxygen contain as many atoms as in
 (a) 10 g of hydrogen
 (b) 5 g of hydrogen
 (c) 80 g of hydrogen
 (d) 1 g of hydrogen (Karnataka C.E.T. 2007)

Answers

B3. (b)	B4. (b)	B5. (a)	B6. (d)	B7. (d)	B8. (e)	B9. (a)	B10. (d)	B11. (d)	B12. (d)
B13. (a)	B14. (a)	B15. (a)	B16. (d)	B17. (c)	B18. (b)				

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