

**Chapter 12 Stoichiometry Review**

**Directions: Show all work!!! No Work Shown, no credit given! Include all units and chemical formulas.**

- Balance the following equation and work the problems below:  $\text{N}_2 + \text{H}_2 \rightarrow \text{NH}_3$ 
  - If 6.52 L of  $\text{H}_2$  react with excess nitrogen what volume of  $\text{NH}_3$  is produced?
  - In order to produce 6.52 L of  $\text{NH}_3$  how many liters of nitrogen are needed?
  - If  $2.35 \times 10^{24}$  molecules of  $\text{NH}_3$  is formed how many grams of hydrogen was used?
- Use the following equation answer these questions:  $\text{Mg} + 2 \text{HNO}_3 \rightarrow \text{Mg}(\text{NO}_3)_2 + \text{H}_2$ 
  - How many grams of magnesium is need to react with 6.28 grams  $\text{HNO}_3$ ?
  - If I start this reaction with 40.5 grams of magnesium and an excess of nitric acid,
    - how many grams of hydrogen gas will I produce?
    - If 1.7 grams of hydrogen is actually produced, what was my percent yield of hydrogen?
- Use the following equation answer these questions:  $\text{NaHCO}_3 \rightarrow \text{NaOH} + \text{CO}_2$ 
  - Find the theoretical yield in grams of  $\text{NaOH}$  if 14.50 grams  $\text{NaHCO}_3$  is decomposed.
  - If 25 liters of carbon dioxide gas is produced in this reaction,
    - how many grams of sodium hydroxide should be produced?
    - If 50 grams of sodium hydroxide are actually produced, what was my percent yield?
- Use the following equation answer these questions:  $2 \text{FePO}_4 + 3 \text{Na}_2\text{SO}_4 \rightarrow \text{Fe}_2(\text{SO}_4)_3 + 2 \text{Na}_3\text{PO}_4$ 
  - How many grams of  $\text{FePO}_4$  is needed to produce  $2.67 \times 10^{25}$   $\text{Na}_3\text{PO}_4$  formula units
  - If I start with 25 g of  $\text{FePO}_4$ 
    - how many grams of  $\text{Fe}_2(\text{SO}_4)_3$  can I make?
    - If 18.5 grams of  $\text{Fe}_2(\text{SO}_4)_3$  are actually made what is the percent yield?
- Use the following equation answer these questions:  $\text{C}_{12}\text{H}_{22}\text{O}_{11} + 12\text{O}_2 \rightarrow 12\text{CO}_2 + 11\text{H}_2\text{O}$ 
  - If 3.45 g  $\text{CO}_2$  is produced how many water molecules are also produced?
  - If there are 10.0 g of  $\text{C}_{12}\text{H}_{22}\text{O}_{11}$  and 10.0 g of oxygen reacting. Which is the limiting reagent?
- Use the following equation to answer these questions:  $2 \text{Na}_{(s)} + 2 \text{H}_2\text{O}_{(l)} \rightarrow 2 \text{NaOH}_{(aq)} + \text{H}_{2(g)}$ 
  - How many liters of water is needed to produce 87.69 L  $\text{H}_2$ ?
  - If 90.0 grams of sodium is dropped into 80.0 g of water, how many liters of hydrogen would be produced? (hint this is a limiting reactant problem)
- For the following equation answer these questions:  $3\text{PbCl}_2 + \text{Al}_2(\text{SO}_4)_3 \rightarrow 2 \text{AlCl}_3 + 3\text{PbSO}_4$ 
  - What mass of  $\text{PbCl}_2$  is needed to produce 128 g of  $\text{PbSO}_4$ ?
  - If 25.5 g of  $\text{Al}_2(\text{SO}_4)_3$  reacts with excess  $\text{PbCl}_2$  how much aluminum chloride in grams is produced?
  - What is the limiting reactant if you have 65.45 g of each reactant?
  - What is the excess reactant? How many grams of excess reactant are left over?
  - What is the theoretical yield in grams of the Aluminum Chloride if you react 85 g of  $\text{PbCl}_2$ ?
    - If the actual yield is 42.5 g, what is the % yield?
- Balance the following equation and work the problem.  $\text{FeS} + \text{HCl} \rightarrow \text{H}_2\text{S} + \text{FeCl}_2$ 
  - What mass of Iron(II) Sulfide ( $\text{FeS}$ ) is needed to produce 42.3 L of  $\text{H}_2\text{S}$ ?
  - If 62.43 g of  $\text{HCl}$  is reacted with excess  $\text{FeS}$  how many grams of  $\text{FeCl}_2$  is produced?
  - What is the limiting reactant if you have 23.50 g of each reactant?
  - What is the excess reactant? How many grams of excess reactant are left over?
  - What is the theoretical yield in grams of the  $\text{FeCl}_2$  if you react 22.35 g  $\text{FeS}$  with excess  $\text{HCl}$ ?
    - If the actual yield is 12.5 g, what is the % yield?

# Chemistry Matter And Change Chapter 11 Study Guide

## Answers Stoichiometry

**E Staff**



## **Chemistry Matter And Change Chapter 11 Study Guide Answers Stoichiometry:**

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