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**SOLUTIONS MANUAL**

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## CHAPTER 2 THE COMPONENTS OF MATTER

### **FOLLOW-UP PROBLEMS**

- 2.1A **Plan:** An element has only one kind of atom; a compound is composed of at least two kinds of atoms. A mixture consists of two or more substances mixed together in the same container.

**Solution:**

(a) There is only one type of atom (blue) present, so this is an **element**.

(b) Two different atoms (brown and green) appear in a fixed ratio of 1/1, so this is a **compound**.

(c) These molecules consist of one type of atom (orange), so this is an **element**.

- 2.1B **Plan:** An element has only one kind of atom; a compound is composed of at least two kinds of atoms.

**Solution:**

The circle on the left contains molecules with either only orange atoms or only blue atoms. This is a **mixture of two different elements**. In the circle on the right, the molecules are composed of one orange atom and one blue atom so this is a **compound**.

- 2.2A **Plan:** Use the mass fraction of uranium in pitchblende (from Sample Problem 2.2) to find the mass of pitchblende that contains 2.3 t of uranium. Subtract the amount of uranium from that amount of pitchblende to obtain the mass of oxygen in that amount of pitchblende. Find the mass fraction of oxygen in pitchblende and multiply the amount of pitchblende by the mass fraction of oxygen to determine the mass of oxygen in the sample.

**Solution:**

$$\text{Mass (t) of pitchblende} = 2.3 \text{ t uranium} \times \frac{84.2 \text{ t pitchblende}}{71.4 \text{ t uranium}} = 2.7123 = 2.7 \text{ t pitchblende}$$

$$\text{Mass (t) of oxygen in 84.2 t of pitchblende} = 84.2 \text{ t pitchblende} - 71.4 \text{ t uranium} = 12.8 \text{ t oxygen}$$

$$\text{Mass (t) of oxygen} = 2.7123 \text{ t pitchblende} \times \frac{12.8 \text{ t oxygen}}{84.2 \text{ t pitchblende}} = 0.4123 = 0.41 \text{ t oxygen}$$

- 2.2B **Plan:** Subtract the amount of silver from the amount of silver bromide to find the mass of bromine in 26.8 g of silver bromide. Use the mass fraction of silver in silver bromide to find the mass of silver in 3.57 g of silver bromide. Use the mass fraction of bromine in silver bromide to find the mass of bromine in 3.57 g of silver bromide.

**Solution:**

$$\text{Mass (g) of bromine in 26.8 g silver bromide} = 26.8 \text{ g silver bromide} - 15.4 \text{ g silver} = 11.4 \text{ g bromine}$$

$$\text{Mass (g) of silver in 3.57 g silver bromide} = 3.57 \text{ g silver bromide} \times \frac{15.4 \text{ g silver}}{26.8 \text{ g silver bromide}} = 2.05 \text{ g silver}$$

# Chemistry Matter And Change Assessment Solution Manual

**Patrick Vollmar**



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