

Final Exam (90 Multiple Choice)

Identify the letter of the choice that best completes the statement or answers the question.

1. The periodic law states that the properties of elements are periodic functions of their atomic numbers. This means that the _____ determines the position of each element in the periodic table.

- A) mass number
B) number of neutrons
C) number of protons
D) number of nucleons

2. Which of the following is the most reactive element?

- A) helium
B) radon
C) francium
D) hydrogen
E) ununoctium

3. Elements in a group or column in the periodic table can be expected to have similar

- A) atomic masses.
B) atomic numbers.
C) numbers of neutrons.
D) properties.

4. Which of the following gives the correct numbers of protons, neutrons, and electrons in a neutral atom of $^{118}_{50}\text{Sn}$?

- A) 118 protons, 50 neutrons, 118 electrons
B) 118 protons, 118 neutrons, 50 electrons
C) 50 protons, 68 neutrons, 50 electrons
D) 68 protons, 68 neutrons, 50 electrons
E) 50 protons, 50 neutrons, 50 electrons

5. Which group shows the correct order of first ionization energy?

- A) Ba > Ca > Be
B) Cl > Al > P
C) F < Cl < Br
D) B > C > N
E) Cs < K < Na

6. If an atom has 4 valence electrons, to which group of elements does it belong?

- A) group headed by carbon
B) group headed by lithium
C) group headed by boron
D) group headed by nitrogen

7. A metallic element X forms a phosphate with the formula $\text{X}(\text{PO}_3)_2$; the formula for the oxide of X is?

- A) X_2O_3
B) XO
C) XO_2
D) X_2O_4
E) X_2O_5

8. According to present theory which of the following compounds would be classified as an ionic compound?

- A) nitrogen trichloride
B) calcium oxide
C) ammonia
D) carbon dioxide
E) methane

Chemistry Ii Final Multiple Choice

Jianjun Gao



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





















































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