

This print-out should have 40 questions. Multiple-choice questions may continue on the next column or page – find all choices before answering.

001 10.0 points

Calculate the number of H_2O molecules in 1.00 cm^3 of water at 0°C (density = 0.99987 g/cm^3).

1. 8.36×10^{24} molec
2. 1.55×10^{23} molec
3. 6.69×10^{22} molec
4. 3.35×10^{22} molec **correct**

Explanation:

002 10.0 points

How many moles of hydrogen are contained in 3.00 moles of ethanol ($\text{CH}_3\text{CH}_2\text{OH}$)?

1. $3.00 \times 6.02 \times 10^{23}$ mol
2. $3.00 \times 3.61 \times 10^{24}$ mol
3. 3.61×10^{24} mol
4. 6.02×10^{23} mol
5. 18.00 mol **correct**
6. 3.00×10^{23} mol
7. 3.00 mol
8. 1.00×10^{23} mol
9. 1.00 mol
10. 6.00 mol

Explanation:

$n = 3 \text{ mol}$

From the molecular formula of ethanol ($\text{CH}_3\text{CH}_2\text{OH}$) we can see that for every molecule of ethanol there are 6 atoms of hydrogen. Therefore, in one mole of ethanol

there would be 6 moles of hydrogen atoms. If this isn't immediately clear, it can be seen in the conversion below using Avogadro's number to convert from atoms and molecules to moles. Starting with our 6 atoms H per 1 molecule ethanol (etOH) ratio:

$$\begin{aligned} ? \frac{\text{mol H}}{\text{mol EtOH}} &= \frac{6 \text{ atoms H}}{1 \text{ molec CH}_3\text{CH}_2\text{OH}} \\ &\times \frac{1 \text{ mol H}}{6.022 \times 10^{23} \text{ atoms H}} \\ &\times \frac{6.022 \times 10^{23} \text{ molec etOH}}{1 \text{ molec etOH}} \\ &= \frac{6 \text{ mol H}}{1 \text{ mol CH}_3\text{CH}_2\text{OH}} \end{aligned}$$

Note, the 6.022×10^{23} factors cancel, leaving us with 6 mol H/1 mol $\text{CH}_3\text{CH}_2\text{OH}$. Thus in 3.00 moles of ethanol, there would be 3.00 times as many H atoms, or 18.00 mol H.

003 10.0 points

Find the molar mass for $(\text{NH}_4)_2\text{CrO}_4$.

1. 168.10 g/mol
2. 110.13 g/mol
3. 136.10 g/mol
4. 142.20 g/mol
5. 152.10 g/mol **correct**

Explanation:

molar mass of $(\text{NH}_4)_2\text{CrO}_4 = ?$

$$\begin{aligned} 2 \text{ mol N} &\times \frac{14.01 \text{ g N}}{\text{mol N}} = 28.02 \text{ g} \\ 8 \text{ mol H} &\times \frac{1.01 \text{ g H}}{\text{mol H}} = 8.08 \text{ g} \\ 1 \text{ mol Cr} &\times \frac{52.00 \text{ g Cr}}{\text{mol Cr}} = 52.00 \text{ g} \\ 4 \text{ mol O} &\times \frac{16.00 \text{ g O}}{\text{mol O}} = 64.00 \text{ g} \end{aligned}$$

$$\text{Molar mass} = 152.10 \text{ g/mol}$$

004 10.0 points

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RM Cervero

A red circular graphic with a gradient, appearing as a stylized arrow or a decorative element, positioned to the right of the light blue bar.

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