## ACID & BASE CHEMISTRY

#### STRONG ACIDS

HCIO<sub>4</sub> HCIO<sub>3</sub> H<sub>2</sub>SO<sub>4</sub> HNO<sub>3</sub> H<sub>3</sub>O+ HI HBr

Percloric Acid
Chloric Acid
Sulfuric Acid
Nitric Acid
Hydrodium ion
Hydroidonic Acid
Hydrobromic Acid

#### WEAK ACIDS

H<sub>3</sub>PO<sub>4</sub> HNO<sub>2</sub> HF CH<sub>3</sub>COOH H<sub>2</sub>CO<sub>3</sub> H<sub>2</sub>S HCOOH H<sub>2</sub>SO<sub>3</sub> Phosphoric Acid Nitrous Acid Hydrofluroic Acid Ethanoic Acid Carbonic Acid Hydrogen Sulfide Methanoic Acid Sulfurous Acid

#### STRONG BASES

NaOH LiOH KOH Ca(OH)<sub>2</sub> Ba(OH)<sub>2</sub> Na<sub>2</sub>CO<sub>3</sub>

Sodium Hydroxide Lithium Hydroxide Potassium Hydroxide Calcium Hydroxide Barium Hydroxide Sodium Carbonate

#### WEAK BASES

NH<sub>3</sub> NaHCO<sub>3</sub>

CH<sub>3</sub>NH<sub>2</sub>

Ammonia Sodium hydrogen carbonate

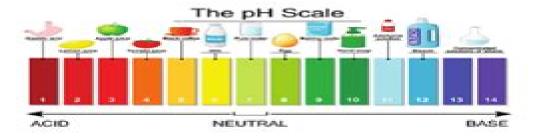
Methylamine

#### TYPES OF NEUTRALISATION REACTIONS

Hydroxide Base + Acid ---> Salt + Water

Carbonate Base + Acid ---> Carbon dioxide + Salt + Water

Ammonia + Acid ---> Ammonium ion + Anion (from Acid)



# **Chemistry Guide Acids And Bases Answe**

**Rodney L. Doran** 

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