# Gas Laws

21. Q: What is the relationship between volume and moles in Avogadro's Principle?

A: Volume is directly proportional to the number of moles.

Q: Calculate the pressure of 1.0 mole of gas in a 10.0 L container at 273
K using the ideal gas law.

A: The pressure is 2.24 atm.

23. Q: According to the ideal gas law, what happens to the pressure if the volume is reduced by half while the temperature and number of moles remain constant?

A: The pressure doubles.

24. Q: If the volume of a gas is tripled and the temperature is doubled, what happens to the pressure according to the combined gas law?

A: The pressure decreases to 2/3 of its original value.

Q: Using gas stoichiometry, how many liters of CO<sub>2</sub> are produced from
Q L of O<sub>2</sub> in the reaction C+O<sub>2</sub>+CO<sub>2</sub> at STP?

A: 5.0 L of CO, are produced.

 Q: In the reaction 2NO+O<sub>2</sub>+2NO<sub>2</sub>, how many liters of NO<sub>2</sub> are produced from 4.0 L of NO at STP?

A: 4.0 L of NO<sub>2</sub> are produced.

27. Q: Calculate the volume of  $N_2$  gas required to produce 20.0 L of  $NH_3$  in the reaction  $N_2 + 3H_2 + 2NH_3$  at STP.

A: 10.0 L of No are required.

28. Q: If the temperature of a gas is increased from 300 K to 600 K and the volume is kept constant, what happens to the pressure?

A: The pressure doubles.

29. Q: Using the ideal gas law, how do you calculate the volume of gas if you know the pressure, temperature, and number of moles?

A: VenRT/P

30. Q: In the reaction H<sub>2</sub>+Cl<sub>2</sub>+2HCl, how many liters of HCl are produced from 3.0 L of H<sub>2</sub> at STP?

A: 5.0 L of HCI are produced.



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