

Chemistry 100 Final Study Guide Key

• Chapter 1: Introduction

a. Physical vs. Chemical properties

1) Physical properties: shiny, metallic, malleable, ductile, conductive. Chemical property: reactive with water.

b. Units of measurement

1) Meters, liters, grams (with any prefixes) are metric system. All others are U.S.

2) 5.68×10^{-7} 46,300,000

c. Density

1) 2.47g/mL

• Chapter 3: Atomic Structure

a. Historical Perspective

1) List the major scientific accomplishment of the following scientists:

(i) Experiments in electrochemistry described ions and the electrical nature of atoms

(ii) Cathode-ray tube experiments led to the proposal of the "Plum pudding" model of atoms

(iii) Oil-drop experiments calculated the mass of electrons

(iv) Gold foil experiments proved atoms had a dense, positively-charged nucleus and empty space

(v) Proposed an atomic model with a nucleus and electrons circling like planets

(vi) Proposed that electrons exist in a "cloud" or probability region around the nucleus

b. Isotopes

1) Isotopes of an element have the same number of protons, but different numbers of neutrons. Regular hydrogen has 1 proton and no neutrons. Isotopes could have 1 proton and 1 neutron (deuterium) or 1 proton and 2 neutrons (tritium).

c. Structure of the atom

1) Protons and neutrons are in the nucleus. Electrons are in regions of space surrounding the nucleus.

2) 3 energy levels, 4 sublevels, 6 orbitals, 11 electrons

3) $1s^2 2s^2 2p^6 3s^2 3p^2$

d. Electron configurations and trends on the Periodic Table

1) See the periodic Table.

• Chapter 5: Chemical bonds

a. Electron-dot structures

1) Draw the electron-dot structure of an oxygen atom.



b. Octet Rule and Ions

1) Mg^{2+} , Cl^- , N^{3-}

c. Ionic compounds

1) Aluminum bromide, calcium sulfide, sodium nitride

d. Covalent compounds

1) Sulfur dioxide, phosphorus trichloride, carbon tetrahydride

e. Polyatomic molecules

1) Nitrate (NO_3^-), chlorate (ClO_3^-), carbonate (CO_3^{2-}), sulfate (SO_4^{2-}), chromate (CrO_4^{2-}), phosphate (PO_4^{3-})

f. Electron-dot formulas for molecules

1) Answer will be given in class.

g. VSEPR Model and Molecular shape



VSEPR shape = pyramidal

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