

Chemistry 338 2nd Semester Final Exam Review Sheet

District Assessment

Element	Symbol	Atomic #	Protons	Neutrons	Electrons	Mass #	Charge
Tungsten	¹⁸⁴ W ⁺² 74	74	74	110	72	184	+2

1. Name the compound
 - a. Fe₂O₃ – iron (III) oxide
 - b. Cl₂O₈ – dichlorine octaoxide

2. Write the formula
 - a. Trisulfur Dinitride - S₃N₂
 - b. Zinc Hydroxide – Zn(OH)₂
 - c. Tin (IV) oxide - SnO₂

3. The pressure of a gas is 843 mm Hg at 24°C. What will the pressure be at standard temperature? Volume is constant.

$$\frac{P_1}{T_1} = \frac{P_2}{T_2} \quad \frac{843\text{mmHg}}{297\text{K}} = \frac{P_2}{273\text{K}} \quad P_2 = 775 \text{ mm Hg}$$

4. The volume of a gas is 81.7 mL when the pressure is 119 kPa and the temperature is 73°C. What is the volume at STP?

$$\frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2} \quad \frac{(119\text{kPa})(81.7\text{mL})}{346\text{K}} = \frac{(101.3\text{kPa})(V_2)}{273\text{K}} \quad V_2 = 75.7 \text{ mL}$$

5. Between Be and Ba, which has a larger ionization energy? Larger atomic radius?

Larger ionization energy – Be
Larger atomic radius - Ba

6. Balance the equation & classify the type of reaction:

$$2\text{Al} + 3 \text{CuSO}_4 \rightarrow 1\text{Al}_2(\text{SO}_4)_3 + 3 \text{Cu} \quad \text{single replacement}$$

7. Predict the products & classify the type of reaction:

$$\text{H}_3\text{PO}_4 + 3 \text{NaOH} \rightarrow 3 \text{H}_2\text{O} + \text{Na}_3\text{PO}_4 \quad \text{double replacement}$$

8. Calculate the Molar Mass: (NH₄)₂CO₃
96 g/mole

9. Convert 2.75 moles of N₂O₅ to molecules

$$\frac{2.75\text{moles}}{1} \times \frac{6.02 \times 10^{23}\text{molecules}}{1\text{mole}} = 1.66 \times 10^{24} \text{ molecules N}_2\text{O}_5$$

Solutions

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A Loxley



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Table of Contents Chemistry Final Review Semester 2

1. Understanding the eBook Chemistry Final Review Semester 2
 - The Rise of Digital Reading Chemistry Final Review Semester 2
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Final Review Semester 2
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Final Review Semester 2
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Final Review Semester 2
 - Personalized Recommendations
 - Chemistry Final Review Semester 2 User Reviews and Ratings
 - Chemistry Final Review Semester 2 and Bestseller Lists
5. Accessing Chemistry Final Review Semester 2 Free and Paid eBooks
 - Chemistry Final Review Semester 2 Public Domain eBooks
 - Chemistry Final Review Semester 2 eBook Subscription Services
 - Chemistry Final Review Semester 2 Budget-Friendly Options
6. Navigating Chemistry Final Review Semester 2 eBook Formats

- ePub, PDF, MOBI, and More
 - Chemistry Final Review Semester 2 Compatibility with Devices
 - Chemistry Final Review Semester 2 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Final Review Semester 2
 - Highlighting and Note-Taking Chemistry Final Review Semester 2
 - Interactive Elements Chemistry Final Review Semester 2
 8. Staying Engaged with Chemistry Final Review Semester 2
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Final Review Semester 2
 9. Balancing eBooks and Physical Books Chemistry Final Review Semester 2
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Final Review Semester 2
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Chemistry Final Review Semester 2
 - Setting Reading Goals Chemistry Final Review Semester 2
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Chemistry Final Review Semester 2
 - Fact-Checking eBook Content of Chemistry Final Review Semester 2
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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