

Chemistry | Semester 2 | Final Exam Study Guide
Unit 9: Compound Stoichiometry

38. A mole is a unit defined to be $= 6.02 \times 10^{23}$ particles
 1 mole = mass

39. The molar mass of an element is the mass of one mole of the substance
 = its atomic mass in grams

40. What information is given to you in the empirical formula?

- elements in comp.
- ratio of elements

41. What do the subscripts represent in a molecular formula?
 how many atoms of each element

42. Define Formula Mass
 mass of 1 mole of an ionic compound

43. Define Molar Mass
 mass of 1 mole of any compound

44. Define Molecular Mass
 mass of one mole of a covalent (molecular) comp.

True/False

45. F One mole of O_2 has a greater mass than 1 mole of Na_2O . 61.98

46. F One mole of H_2 has ^{same} molecules than 1 mole of Na_2O . same # molecules

47. F One molecule of H_2 has more atoms than 1 molecule of Na_2O . 3 atoms

Calculations

48. Find the molar mass of $Al_2(SO_4)_3$

Find mass of each element \rightarrow multiply each mass by subscript \rightarrow add masses of all elements together = molar mass

Al (2) $26.98 = 53.96$
 S (3) $32.07 = 96.21$
 O (12) $16.00 = 192.00$
294.17 g/mol

49. How many moles are in 2.4g of carbon dioxide (CO_2)?

Start with the given value \rightarrow use dimensional analysis so you can factor out labels until you get to target units

$2.4g CO_2 \times \frac{1 \text{ mol } CO_2}{44 \text{ g } CO_2} = 0.05 \text{ mol}$
C 12 = 12
O (2) 16 = 32
44
 molar mass CO_2

50. How many molecules are contained in 35.37g of water (molar mass of water is 18.02g/mol)?

$35.37g H_2O \times \frac{1 \text{ mol } H_2O}{18.02g H_2O} \times \frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mol } H_2O} = 1.18 \times 10^{24} \text{ molecules } H_2O$

51. How many atoms are in 6.3 moles of calcium?

Hint: there are 6.02×10^{23} atoms in 1 mole

$6.3 \text{ mol Ca} \times \frac{6.02 \times 10^{23} \text{ atoms}}{1 \text{ mol Ca}} = 3.79 \times 10^{24} \text{ atoms}$

52. Calculate the number of moles there are in 4.5×10^{28} molecules of sulfur dioxide (SO_2).

$4.5 \times 10^{28} \text{ molecules } SO_2 \times \frac{1 \text{ mol } SO_2}{6.02 \times 10^{23} \text{ molecules}} = 747.51 \text{ mol } SO_2$

53. Calculate the percent composition of sodium and oxygen in sodium oxide (Na_2O).

Find the molar mass \rightarrow divide total mass of element where it's solving for by the total mass of compound

Na (2) $22.99 = 45.98 / 61.98 = 74.2\%$
 O (1) $16.00 = 16.00 / 61.98 = 25.8\%$

molar mass

45.98
 16.00
 61.98 g/mol

54. What is the empirical formula of a compound that contains 81.68% carbon and 18.32% hydrogen?

Assume 100 g: % to mass \rightarrow mass to mole \rightarrow divide by smallest \rightarrow multiply till whole

$81.68g C \times \frac{1 \text{ mol}}{12.01g C} = 6.8 / 6.8 = 1 \times 3 = 3$
 $18.32g H \times \frac{1 \text{ mol}}{1.01g H} = 18.139 / 6.8 = 2.67 \times 3 = 8$
 C_3H_8

55. What is the molecular formula of a hydrocarbon that has an empirical formula of CH_2 and a molecular mass (molar mass of molecule) of 48.15 g/mol?

$\frac{MM_{\text{molecule}}}{MM_{\text{empirical}}} = \frac{48.15}{16.05} = 3$
 C_3H_6

Multiply each subscript by the above answer to get the subscripts for the molecular formula

To find Empirical Formula
 1. Divide (multiply till whole)

if it ends in multiply by
 0.2 or 0.8 $\times 5$
 0.25 or 0.75 $\times 4$
 0.33 or 0.67 $\times 3$
 0.5 $\times 2$

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1. Understanding the eBook Chemistry Final Review Guide Answers
 - The Rise of Digital Reading Chemistry Final Review Guide Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Final Review Guide Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Final Review Guide Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Final Review Guide Answers
 - Personalized Recommendations
 - Chemistry Final Review Guide Answers User Reviews and Ratings
 - Chemistry Final Review Guide Answers and Bestseller Lists
5. Accessing Chemistry Final Review Guide Answers Free and Paid eBooks
 - Chemistry Final Review Guide Answers Public Domain eBooks
 - Chemistry Final Review Guide Answers eBook Subscription Services
 - Chemistry Final Review Guide Answers Budget-Friendly Options

6. Navigating Chemistry Final Review Guide Answers eBook Formats
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 - Chemistry Final Review Guide Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Final Review Guide Answers
 - Highlighting and Note-Taking Chemistry Final Review Guide Answers
 - Interactive Elements Chemistry Final Review Guide Answers
8. Staying Engaged with Chemistry Final Review Guide Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Final Review Guide Answers
9. Balancing eBooks and Physical Books Chemistry Final Review Guide Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Final Review Guide Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Final Review Guide Answers
 - Setting Reading Goals Chemistry Final Review Guide Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Final Review Guide Answers
 - Fact-Checking eBook Content of Chemistry Final Review Guide Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

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